

## IONIC COVALENT BONDING WEBQUEST

### IONIC BONDING

WEBSITE 1: <http://visionlearning.com/en/library/Chemistry/1/Chemical-Bonding/55>

1. Approximately how many elements are represented on the periodic table?  
118
2. What accounts for the fact that there are far more substances than are listed on the periodic table?  
there are more substances because elements can combine and react to form new substances
3. "Formed when two or more atoms chemically bond together, the resulting compound is unique both chemically and physically from its parent atoms."
4. What compound forms during the reaction between elemental sodium and elemental chlorine?  
Name: sodium chloride  
Formula: NaCl
5. List four facts related to G.N. Lewis:  
  - atoms with fewer than eight electrons bond together
  - he is an american chemist
6. Explain the behavior of electrons in ionic bonding.  
electrons are transferred from one atom to another
7. In the process of either gaining or losing negatively charged electrons, the reacting atoms form ions.
8. In the reaction between sodium and chlorine, which atom loses an electron?  
Na loses e<sup>-</sup> Which atom gains an electron? Cl gains e<sup>-</sup>
9. After transferring the electron, which ion is negatively charged? The anion  
Why is the ion negatively charged? it gains an electron Write the symbol for the negative ion: Anion(-)
10. After transferring the electron, which ion is positively charged? the cation ion  
Why is the ion positively charged? it loses electrons  
Write the symbol for the positive ion: C(+)
11. How does the sodium atom contrast to the sodium ion? (charge, size)  
The sodium loses the electrons it becomes smaller in size and its charge become positive.

# Ionic Covalent Bonding Webquest Answers Key Page

**Marcel A. Müller**



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