

WEEK 1 & 2

March 2023, sept 2023

1.

a. The load details of an apartment are as follows:

- Total number of houses = 16
- Connected load of each house = 5 KW
- Emergency lighting for each house = 100 W
- Common area lighting = 60 Nos. of 10 W LED fitting
- Common area raw power = 2 KW
- Borewell pump 3 phase 3 HP = 1 No and 3 phase 2 HP = 1 No.
- Lift 5 HP = 1 no.

Determine the total connected load. From the above loads, identify the loads which are very during power cuts and select a suitable DG set for the same

Solution

Calculation of Total Connected Load

1. Total Connected Load of Houses

- Total number of houses=16

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Huangqi Zhang



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monoxide co determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2 was produced i what mass of iodine was produced

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 $qquad Cu + 2 AgNO_3 \rightarrow 2 Ag + Cu(NO_3)_2$

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 $3CO + 7H_2 \rightarrow C_3H_8 + 3H_2O$

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation
 $8.3 \text{ mole } H_2S \times \frac{2 \text{ mole } H_2O}{2 \text{ mole } H_2S} = 8.3 \text{ mole } H_2O$
8.3 mole $H_2S \times 18 \text{ g } H_2O / \text{mole } H_2O = 149.4 \text{ g } H_2O$ is your theoretical yield so percent yield
 $\frac{137.1 \text{ g } H_2O}{149.4 \text{ g } H_2O} \times 100 = 91.77\%$
problems 6 8 are similar to problem 5

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 $C_3H_6 + 3O_2 \rightarrow 3CO + 3H_2O$
if the reaction of 21.3 grams of O_2 produces a 47.3% yield how many grams of co would be produced step one figure out how many moles of the limiting reagent

you have 0.216232 grams mole 21.3 grams 0.66 moles 32 g m

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained

calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al 3 O_2 2 Al_2O_3 b determine the theoretical yield of Al_2O_3 160.0

g Al $\times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield

yield actual $\times 100$ 260.0 g $\times 100$ 86.01 theoretical 302.3 g

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web 7 Na 2 C 2 O 4 is the limiting reactant percent yield 86.6 8 only four molecules can be made 9 this amount cannot be

weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12

C is the limiting reagent 4 33 g of H 2 are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s

remains when the reaction is complete

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web the percent yield is calculated as follows text percent yield $\frac{\text{text actual yield}}{\text{text theoretical yield}} \times 100$ percent

yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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web may 20 2018 learning objectives identify the limiting reactant limiting reagent in a given chemical reaction calculate

how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is

complete

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web limiting reagent and percent yield practice worksheet answer key limiting reactant yield practice worksheet 1 methanol

CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is

reacted with 8.6 g of H_2

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problems from a balanced chemical equation 3 sto 5 differentiate between the actual yield and theoretical yield of a chemical

reaction

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web limiting reagents and percentage yield worksheet 1 consider the reaction $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow \text{I}_2 + 5\text{CO}_2$ a 80.0 grams of iodine v oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

7 3 limiting reactant and percent yield problems - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3