

- (e) Oxygen is more electronegative than carbon and attracts electrons from it. Formaldehyde has a dipole moment.



- (f) Nitrogen is more electronegative than carbon. Hydrogen cyanide has a dipole moment.



- 1.20** The orbital diagram for sp^3 -hybridized nitrogen is the same as for sp^3 -hybridized carbon, except nitrogen has one more electron.



The unshared electron pair in ammonia (NH_3) occupies an sp^3 -hybridized orbital of nitrogen. Each N—H bond corresponds to overlap of a half-filled sp^3 hybrid orbital of nitrogen and a 1s orbital of hydrogen.

- 1.21** Silicon lies below carbon in the periodic table, and it is reasonable to assume that both carbon and silicon are sp^3 -hybridized in H_3CSiH_3 . The C—Si bond and all of the C—H and Si—H bonds are σ bonds.



The principal quantum number of the carbon orbitals that are hybridized is 2; the principal quantum number for the silicon orbitals is 3.

- 1.22** (b) Carbon in formaldehyde ($\text{H}_2\text{C}=\text{O}$) is directly bonded to three other atoms (two hydrogens and one oxygen). It is sp^2 -hybridized.
 (c) Ketene has two carbons in different hybridization states. One is sp^2 -hybridized; the other is sp -hybridized.



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