

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 12****STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of  $\text{O}_2$ ,  $\text{N}_2$ , and  $\text{NO}_2$  are 32.00 g/mol, 28.02 g/mol, and 46.01 g/mol, respectively.

Amount of $\text{O}_2$	Amount of $\text{N}_2$	Amount of $\text{NO}_2$	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	$\text{N}_2$	2 molecules $\text{O}_2$
2 molecules	4 molecules	4 molecules	$\text{O}_2$	2 molecules $\text{NO}_2$
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	4.00 mol	4.00 mol	$\text{N}_2$	2.00 mol $\text{O}_2$
2.00 mol	4.00 mol	4.00 mol	$\text{O}_2$	2.00 mol $\text{NO}_2$
1.00 mol	4.00 mol	4.00 mol	$\text{O}_2$	2.00 mol $\text{NO}_2$
0.500 mol	0.200 mol	0.200 mol	$\text{N}_2$	0.300 mol $\text{O}_2$
32.00 g	56.04 g	92.02 g	none	none
64.00 g	56.04 g	92.02 g	$\text{O}_2$	32.00 g $\text{NO}_2$
64.00 g	112.08 g	184.04 g	$\text{O}_2$	1.24 g $\text{NO}_2$

balanced equation mole ratio = 2 mol  $\text{NO}_2$ /1 mol  $\text{O}_2$

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.5125 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{N}_2 \times 1 \text{ mol } \text{N}_2 / 28.02 \text{ g } \text{N}_2 = 0.7138 \text{ mol } \text{N}_2$

actual mole ratio =  $0.7138 \text{ mol } \text{N}_2 / 0.5125 \text{ mol } \text{O}_2 = 1.392 \text{ mol } \text{N}_2 / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of  $\text{N}_2/\text{O}_2$  is larger than the balanced equation mole ratio of  $\text{N}_2/\text{O}_2$ , there is an excess of  $\text{N}_2$ ;  $\text{O}_2$  is the limiting reactant.

Mass of  $\text{NO}_2$  used =  $0.5125 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO}_2 / 1 \text{ mol } \text{O}_2 = 0.4250 \text{ mol } \text{NO}_2$

$0.4250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 19.56 \text{ g } \text{NO}_2$

Mass of  $\text{NO}_2$  produced =  $0.4250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 19.56 \text{ g } \text{NO}_2$

Excess  $\text{N}_2$  =  $20.00 \text{ g } \text{N}_2 - 19.56 \text{ g } \text{N}_2 = 0.44 \text{ g } \text{N}_2$

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**CHAPTER 13****STUDY GUIDE FOR CONTENT MASTERY****Section 13.4 Percent Yield**

In your textbook, read about the yields of products.

Study the diagram and the example problem.



**Example Problem:** The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and  $\text{Ga}_2\text{O}_3$  are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

1. actual yield unknown 7.42 g  $\text{Ga}_2\text{O}_3$
2. mass of reactant: 7.00 g Ga correct
3. number of moles of reactant:  $7.00 \text{ g Ga} \times \frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}} = 7.00 \text{ g Ga} \times 1 \text{ mol Ga} / 69.72 \text{ g Ga}$
4. number of moles of product:  $0.100 \text{ mol Ga} \times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol Ga}} = 0.200 \text{ mol Ga} \times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol Ga}}$
5. theoretical yield:  $0.200 \text{ mol } (\text{Ga}_2\text{O}_3) \times \frac{157.44 \text{ g } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } (\text{Ga}_2\text{O}_3)} = 31.49 \text{ g } (\text{Ga}_2\text{O}_3)$  correct
6. percent yield:  $\frac{7.42 \text{ g } (\text{Ga}_2\text{O}_3)}{31.49 \text{ g } (\text{Ga}_2\text{O}_3)} \times 100 = 23.56\%$  correct

# Mcgraw Hill Companies Night Study Guide

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## **Mcgraw Hill Companies Night Study Guide:**

Mcgraw Hill Companies Night Study Guide: Bestsellers in 2023 The year 2023 has witnessed a remarkable surge in literary brilliance, with numerous compelling novels captivating the hearts of readers worldwide. Lets delve into the realm of top-selling books, exploring the fascinating narratives that have enthralled audiences this year. The Must-Read : Colleen Hoover's "It Ends with Us" This heartfelt tale of love, loss, and resilience has captivated readers with its raw and emotional exploration of domestic abuse. Hoover masterfully weaves a story of hope and healing, reminding us that even in the darkest of times, the human spirit can triumph. Uncover the Best : Taylor Jenkins Reids "The Seven Husbands of Evelyn Hugo" This intriguing historical fiction novel unravels the life of Evelyn Hugo, a Hollywood icon who defies expectations and societal norms to pursue her dreams. Reids compelling storytelling and compelling characters transport readers to a bygone era, immersing them in a world of glamour, ambition, and self-discovery. Mcgraw Hill Companies Night Study Guide : Delia Owens "Where the Crawdads Sing" This captivating coming-of-age story follows Kya Clark, a young woman who grows up alone in the marshes of North Carolina. Owens spins a tale of resilience, survival, and the transformative power of nature, captivating readers with its evocative prose and mesmerizing setting. These popular novels represent just a fraction of the literary treasures that have emerged in 2023. Whether you seek tales of romance, adventure, or personal growth, the world of literature offers an abundance of engaging stories waiting to be discovered. The novel begins with Richard Papen, a bright but troubled young man, arriving at Hampden College. Richard is immediately drawn to the group of students who call themselves the Classics Club. The club is led by Henry Winter, a brilliant and charismatic young man. Henry is obsessed with Greek mythology and philosophy, and he quickly draws Richard into his world. The other members of the Classics Club are equally as fascinating. Bunny Corcoran is a wealthy and spoiled young man who is always looking for a good time. Charles Tavis is a quiet and reserved young man who is deeply in love with Henry. Camilla Macaulay is a beautiful and intelligent young woman who is drawn to the power and danger of the Classics Club. The students are all deeply in love with Morrow, and they are willing to do anything to please him. Morrow is a complex and mysterious figure, and he seems to be manipulating the students for his own purposes. As the students become more involved with Morrow, they begin to commit increasingly dangerous acts. The Secret History is a exceptional and gripping novel that will keep you speculating until the very end. The novel is a cautionary tale about the dangers of obsession and the power of evil.

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