

5 Descriptions of the bonding in three substances are given.

- substance 1 strong covalent bonds between atoms, permanent dipole-dipole attractions between molecules
- substance 2 strong covalent bonds between atoms, weak forces between molecules
- substance 3 strong covalent bonds between atoms, hydrogen bonding between molecules

Which compounds could be substances 1, 2 and 3?

	substance 1	substance 2	substance 3
A	CH ₃ OH	Al ₂ Cl ₆	CH ₂ Cl ₂
B	Al ₂ Cl ₆	CH ₂ Cl ₂	CH ₄
C	CH ₂ Cl ₂	CH ₄	CH ₃ OH
D	CH ₄	CH ₃ OH	H ₂ O

6 Which type of bonding is **never** found in elements?

- A** covalent
- B** ionic
- C** metallic
- D** van der Waals' forces

7 Using the information in the table, what is the enthalpy change, q , for the reaction described?



standard enthalpy change	value / kJ mol ⁻¹
$\Delta H_{\text{f}}^\circ \text{ CsBr}(\text{s})$	+25.9
$\Delta H_{\text{f}}^\circ \text{ Cs}^+(\text{g})$	-276
$\Delta H_{\text{f}}^\circ \text{ Br}^-(\text{g})$	-335

- A** -636.9 **B** -585.1 **C** +585.1 **D** +636.9

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Azhar ul Haque Sario



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