



- 1 The electrolysis of molten aluminium fluoride produces aluminium (Al) at the negative electrode and fluorine (F₂) at the positive electrode.
- a Explain why solid aluminium fluoride does not conduct electricity but molten aluminium fluoride does.
as a solid the ions cannot move to carry charge
but when molten the ions can move to carry charge
 - b Give the formula of the ions in aluminium fluoride: aluminium ions Al^{3+} fluoride ions F^{-}
 - c Give the formula of aluminium fluoride: AlF_3
 - d Explain why aluminium ions go to the negative electrode.
aluminium ions are positive and they are attracted to the negative electrode as opposite charges attract
 - e Write a half equation for the formation of aluminium at the negative electrode. $\text{Al}^{3+} + 3\text{e}^{-} \rightarrow \text{Al}$
 - f Explain whether the formation of aluminium is an oxidation or reduction process.
reduction as the aluminium ions gain electrons
 - g Explain whether the negative electrode is the anode or cathode.
cathode as this is where reduction occurs
 - h Write a half equation for the formation of fluorine at the positive electrode. $2\text{F}^{-} - 2\text{e}^{-} \rightarrow \text{F}_2$
- 2 Complete the table about the electrolysis of some molten ionic compounds.

molten substance	formula	negative electrode half-equation	positive electrode half-equation
calcium oxide	CaO	$\text{Ca}^{2+} + 2\text{e}^{-} \rightarrow \text{Ca}$	$2\text{O}^{2-} - 4\text{e}^{-} \rightarrow \text{O}_2$
magnesium bromide	MgBr_2	$\text{Mg}^{2+} + 2\text{e}^{-} \rightarrow \text{Mg}$	$2\text{Br}^{-} - 2\text{e}^{-} \rightarrow \text{Br}_2$
sodium chloride	NaCl	$\text{Na}^{+} + \text{e}^{-} \rightarrow \text{Na}$	$2\text{Cl}^{-} - 2\text{e}^{-} \rightarrow \text{Cl}_2$

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