

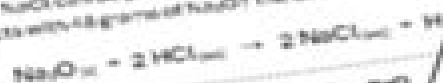
# Stoichiometry

## ● Limiting Reactant & Percentage Yield ●

### Answers

#### Stoichiometry (Page 8) Level 5: Limiting Reactants

- 5) How many grams of NaCl can be produced from the reaction when 50 g solution of HCl reacts with 48 grams of NaOH? The chemical equation is



Mole of NaOH	Mole	NaOH	HCl	NaCl
$n = \frac{m}{M}$	$n = 0.775 \text{ mol}$	48	5.275	1.55
Mole of HCl	$n = 0.905 \text{ mol}$	48	5.275	1.55
$n = \frac{m}{M}$	$n = 0.905 \text{ mol}$	48	5.275	1.55

- 6) How many liters of  $\text{CO}_2$  at STP can be produced when 50 g of benzene ( $\text{C}_6\text{H}_6$ ) reacts with 50 g of  $\text{O}_2$ ?



Mole of $\text{C}_6\text{H}_6$	Mole	$\text{C}_6\text{H}_6$	$\text{O}_2$
$n = \frac{m}{M}$	$n = 0.375 \text{ mol}$	50	1.56
Mole of $\text{O}_2$	$n = 1.56 \text{ mol}$	50	1.56

### Answers

#### Stoichiometry (Page 12) Level 6: Percentage Yield

1. If 15.8 grams of methane ( $\text{CH}_4$ ) undergo complete combustion in excess oxygen to produce 29.4 grams of water, what is the percentage yield of this reaction? The chemical equation is balanced.



$n = \frac{m}{M}$	$n = 0.985 \text{ mol}$	$\text{CH}_4 + \text{H}_2\text{O}$	$1 : 2$	$n = \frac{m}{M}$
$n = \frac{m}{M}$	$n = 1.56 \text{ mol}$	$0.985 : 1$		$2.97 = \frac{m}{18.02}$
$n = 0.985 \text{ mol}$		$n = 1.97 \text{ mol}$		$m = 35.5 \text{ g/mol}$
Percentage (%) Yield				
$\% = \frac{\text{actual}}{\text{theoretical}} \times 100$				

#### Percentage (%) Yield

$$\% = \frac{\text{Actual}}{\text{Theoretical}} \times 100 = \frac{29.4}{35.3} \times 100 = 83.3\%$$

Answer: **83.3%**

2. If 78.3 grams of sodium carbonate ( $\text{Na}_2\text{CO}_3$ ) reacts with 833 mL of 1.12 M solution of hydrochloric acid to produce 49.7 grams of sodium chloride, what is the limiting reactant and the percentage yield of this reaction? The chemical equation is unbalanced.



Mole of $\text{Na}_2\text{CO}_3$	Mole	$\text{Na}_2\text{CO}_3$	HCl	NaCl	$\text{CO}_2$	$\text{H}_2\text{O}$
$n = \frac{m}{M}$	$n = 0.375 \text{ mol}$	78.3	1.75	3.5	0.375	0.375
Mole of HCl	$n = 0.905 \text{ mol}$	78.3	1.75	3.5	0.375	0.375

Answer Key available

Limiting Reactant: **Hydrochloric acid**

# Limiting Reactants And Percentage Yield Quiz

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