

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

chemistry

LIMITING REACTANTS

1 The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
 If you start with 10 moles of N_2 and 30 moles of H_2 , determine the limiting reactant.

2 Calculate the grams of ammonia produced.

3 How many grams of the excess reactant is produced in a left over?
 To multiply the moles of the limiting reactant by the mole ratio of the excess reactant.

4 Determine moles of product & mass of product.

5 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 60 grams of oxygen, determine the limiting reactant.

6 How many grams of water (steam) are produced?

7 Determine moles of product & mass of product.



8 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure iron ore is produced.
 $Mg + 2FeCl_2 \rightarrow MgCl_2 + 2Fe$
 If you start with 10 moles of magnesium and 2 moles of iron(II) chloride, determine the limiting reactant.

9 Calculate the grams of magnesium chloride produced.

10 How many grams of the excess reactant is produced in a left over?

11 Magnesium chloride is used to create salts for de-icing roads. The reaction with very impure iron ore is produced.
 $Mg + 2FeCl_2 \rightarrow MgCl_2 + 2Fe$
 If you have with 100 grams of magnesium and 60 grams of iron(II) chloride, determine the limiting reactant.

12 How much carbon dioxide is produced?

13 Determine moles of product & mass of product.

12 The reaction below produces carbon dioxide and water from glucose (sugar) and oxygen.
 $C_6H_{12}O_6 + 6O_2 \rightarrow 6CO_2 + 6H_2O$
 If you have 100 grams of glucose and 60 grams of oxygen, determine the limiting reactant.

13 Calculate the grams of carbon dioxide produced.
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 148 \text{ g } CO_2$
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } CO_2}{6 \text{ mol } O_2} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 66 \text{ g } CO_2$
 Since 66 g < 148 g, O_2 is the limiting reactant.

14 How many grams of water (steam) are produced?
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } H_2O}{6 \text{ mol } O_2} \times \frac{18 \text{ g } H_2O}{1 \text{ mol } H_2O} = 165 \text{ g } H_2O$
 Since 66 g < 148 g, O_2 is the limiting reactant.

15 How many moles of carbon dioxide is produced?
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } CO_2}{6 \text{ mol } O_2} = 1.875 \text{ mol } CO_2$
 Since 66 g < 148 g, O_2 is the limiting reactant.

16 How many grams of the excess reactant is produced in a left over?
 $100 \text{ g } C_6H_{12}O_6 - 100 \text{ g } C_6H_{12}O_6 \text{ used} = 0 \text{ g } C_6H_{12}O_6 \text{ left over}$
 Since 66 g < 148 g, O_2 is the limiting reactant.

17 Determine moles of product & mass of product.
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} = 3.33 \text{ mol } CO_2$
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } CO_2}{6 \text{ mol } O_2} = 1.875 \text{ mol } CO_2$
 Since 1.875 mol < 3.33 mol, O_2 is the limiting reactant.

18 Calculate the grams of carbon dioxide produced.
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 148 \text{ g } CO_2$
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } CO_2}{6 \text{ mol } O_2} \times \frac{44 \text{ g } CO_2}{1 \text{ mol } CO_2} = 66 \text{ g } CO_2$
 Since 66 g < 148 g, O_2 is the limiting reactant.

19 How many grams of water (steam) are produced?
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } H_2O}{6 \text{ mol } O_2} \times \frac{18 \text{ g } H_2O}{1 \text{ mol } H_2O} = 165 \text{ g } H_2O$
 Since 66 g < 148 g, O_2 is the limiting reactant.

20 Determine moles of product & mass of product.
 $100 \text{ g } C_6H_{12}O_6 \times \frac{1 \text{ mol } C_6H_{12}O_6}{180 \text{ g } C_6H_{12}O_6} \times \frac{6 \text{ mol } CO_2}{1 \text{ mol } C_6H_{12}O_6} = 3.33 \text{ mol } CO_2$
 $60 \text{ g } O_2 \times \frac{1 \text{ mol } O_2}{32 \text{ g } O_2} \times \frac{6 \text{ mol } CO_2}{6 \text{ mol } O_2} = 1.875 \text{ mol } CO_2$
 Since 1.875 mol < 3.33 mol, O_2 is the limiting reactant.

ANSWER KEY INCLUDED

Limiting Reactant Answer Key

Brendan G. Carr

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Limiting Reactant Answer Key:

Unveiling the Energy of Verbal Beauty: An Emotional Sojourn through **Limiting Reactant Answer Key**

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