

Limiting Reagents and Percentage Yield Worksheet

1. Consider the reaction

*iodine
pentoxide*



a) 80.0 grams of iodine(V) oxide, I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO .

Determine the mass of iodine I_2 , which could be produced?

b) If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.

i) what mass of iodine was produced?

ii) what percentage yield of iodine was produced.

2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

a) Which chemical is the limiting reactant?

b) How many grams of ZnS will be formed?

c) How many grams of the excess reactant will remain after the reaction is over?

3. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen?

What mass is in excess? What mass of MgO is formed?

4. How many grams of Al_2S_3 are formed when 5.00 grams of Al is heated with 10.0 grams S ?

5. When MoO_3 and Zn are heated together they react



What mass of ZnO is formed when 20.0 grams of MoO_3 is reacted with 10.0 grams of Zn ?

6. Silver nitrate, AgNO_3 , reacts with ferric chloride, FeCl_3 , to give silver chloride, AgCl , and ferric nitrate, $\text{Fe}(\text{NO}_3)_3$. In a particular experiment, it was planned to mix a solution containing 25.0 g of AgNO_3 with another solution containing 45.0 grams of FeCl_3 .

a) Write the chemical equation for the reaction.

b) Which reactant is the limiting reactant?

c) What is the maximum number of moles of AgCl that could be obtained from this mixture?

Limiting Reagent And Percent Yield Study Guide

Scott C. Dulebohn



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