2. Limiting and Excess Reactants



The **limiting reactant** is defined as the species in a chemical reaction that **theoretically** would be the first to be completely consumed if the reaction were to proceed to completion according to the chemical equation.

- even if the reaction does not proceed to completion!

All of the other reactant/s are called excess reactants.

For Example:

$$C_7H_{16} + 11O_2 \rightarrow 7CO_2 + 8H_2O$$

If 1 g mol of C_7H_{16} and 12 g mol of O_2 are mixed so as to react,

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Lifeliqe

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<u>Limiting Reactant and Percent</u> Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

Uncover the mysteries within is enigmatic creation, Embark on a Mystery with **Limiting And Excess Reactants Pogil**. This downloadable ebook, shrouded in suspense, is available in a PDF format (*). Dive into a world of uncertainty and anticipation. Download now to unravel the secrets hidden within the pages.

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