

## Limiting Reactants Practice KEY

1. How many grams of aluminum chloride could be produced from 35.5 g of aluminum and 39.0 g of chlorine gas?



$$35.5 \text{ g Al} \times \frac{1 \text{ mol Al}}{26.982 \text{ g Al}} \times \frac{3 \text{ mol Cl}_2}{2 \text{ mol Al}} \times \frac{70.906 \text{ g Cl}_2}{1 \text{ mol Cl}_2} = 140. \text{ g Cl}_2$$

- ∞ This means 35.5 g Al would react with 140. g Cl<sub>2</sub>
- ∞ Only 39.0 g Cl<sub>2</sub> were available, so it is the limiting reactant.
- ∞ Use the mass of Cl<sub>2</sub> in the reaction to determine the mass of the AlCl<sub>3</sub> formed.

$$39.0 \text{ g Cl}_2 = \text{ ? } \text{ g AlCl}_3$$

$$39.0 \text{ g Cl}_2 \times \frac{1 \text{ mol Cl}_2}{70.906 \text{ g Cl}_2} \times \frac{2 \text{ mol AlCl}_3}{3 \text{ mol Cl}_2} \times \frac{133.341 \text{ g AlCl}_3}{1 \text{ mol AlCl}_3} =$$

Answer: 48.9 g AlCl<sub>3</sub>

2. Identify the limiting reagent when 10.8 g water reacts with 4.7 g Na to produce NaOH and H<sub>2</sub>.

To solve, set one reactant to the other and compare what is calculated to the actual amount in the reaction to determine the theoretical limiting reactant.

$$10.8 \text{ g H}_2\text{O} = \text{ ? } \text{ g Na}$$

Or

$$4.7 \text{ g Na} = \text{ ? } \text{ g H}_2\text{O}$$



$$10.8 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.015 \text{ g H}_2\text{O}} \times \frac{2 \text{ mol Na}}{2 \text{ mol H}_2\text{O}} \times \frac{22.990 \text{ g Na}}{1 \text{ mol Na}} = 13.8 \text{ g Na}$$

- ∞ This means that 10.8 g will react completely with 13.8 g Na, but only 4.7 g Na was available to this reaction. Therefore, Na is the limiting reactant in this problem.

# Limiting Reactant Problems Answer Key

**Howard Stephen Stoker**



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*Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

Mapping College Chemistry Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career

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*Investigating Chemistry* Matthew Johl, 2008-12-22 In its new second edition Investigating Chemistry A Forensic Science Perspective remains the only book that uses the inherently fascinating topics of crime and criminal investigations as a context for teaching the

fundamental chemical concepts most often covered in an introductory nonmajors course Covering all the standard topics Matthew Johll capitalizes on the surge of interest in the scientific investigation of crime as sparked by CSI and other television shows bringing together the theme of forensic science and the fundamentals of chemistry in ways that are effective and accessible for students This edition features refined explanations of the chemical concepts which are the core of the book as well as a more thoroughly integrated forensic theme updated features and an expanded media supplements package

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