#### Limiting Reagent Worksheet #1

1. Given the following reaction: (Balance the equation first!)

$$C_3H_8 + 5O_2 -----> 3CO_2 + 4H_2O$$

a) If you start with 14.8 g of C3Ha and 3.44 g of O3, determine the limiting reagent

14.8 g C<sub>3</sub>H<sub>8</sub> x 
$$\frac{1 \mod C3H8}{44.09 g \mod 8}$$
 x  $\frac{4 \mod H2O}{1 \mod C3H8}$  x  $\frac{18.02 g H2O}{1 \mod H2O}$  = 24.2 g H<sub>2</sub>O

3.44 g 
$$O_2$$
 x  $\frac{1 \mod O2}{32gO2}$  x  $\frac{4 \mod H2O}{5 \mod O2}$  x  $\frac{18.02 \text{ g } H2O}{1 \mod H2O}$  = 1.55 g  $H_2O$ 

O, is limiting

b) determine the number of moles of carbon dioxide produced

3.44 g O<sub>2</sub> x 
$$\frac{1 \mod O2}{32 g O2}$$
 x  $\frac{3 \mod CO2}{5 \mod O2}$  x  $\frac{44.01 \text{ g } CO2}{1 \mod CO2}$  = 2.84 g H<sub>2</sub>O

- c) determine the number of grams of H,O produced
- 1.55 g H2O
- d) determine the number of grams of excess reagent left

3.44 g O<sub>2</sub> x 
$$\frac{1 \mod O2}{32gO2}$$
 x  $\frac{1 \mod C3H8}{5 \mod O2}$  x  $\frac{44.09g C3H8}{1 \mod C3H8}$  = .0948 g C<sub>3</sub>H<sub>8</sub> used

$$14.8g C_3H_8 - 0.948g C_3H_8 = 13.9g C_3H_8$$
 remain

2. Given the following equation:

a) If 10.0 g of Al<sub>2</sub>(SO<sub>3</sub>)<sub>3</sub> is reacted with 10.0 g of NaOH, determine the limiting reagent

$$10.00 \text{ g Al}_2(\text{SO}_3)_3 \text{ x } \frac{1 \mod AO(SO3)3}{294.17g AO(SO3)3} \text{ x } \frac{3 \mod No2SO4}{1 \mod AO(SO3)3} \text{ x } \frac{126.05 \text{ g No2SO4}}{1 \mod No2SO4} = 12.9 \text{ g Na}_2\text{SO}_4$$

$$10.00 \text{ g Na}_2 \text{SO}_4 \text{ x } \frac{1 \mod NoOH}{294.17 g \ NoOH} \text{ x } \frac{3 \mod NoOH}{6 \mod NoOH} \text{ x } \frac{126.05 \ g \ No2SO4}{1 \mod NoOH} = 15.8 g \ \text{Na}_2 \text{SO}_4$$

## **Limiting Reagent 1 Answer Key**

**Yicheng Fang** 

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