

Annotated Solution

2018 USNCO Local Exam

Authors: Ritvik Tengamrupa and Harry Dahi
Proofreader and Editor: Dr. Qiang "Steven" Chen

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1 Solutions

1. We can use Avogadro's constant, and then do some dimensional analysis.

$$225 \text{ g O}_2 = \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \times \frac{2 \text{ mol O}}{1 \text{ mol O}_2} = \frac{6.02 \times 10^{23} \text{ atoms}}{\text{mol O}} = \boxed{6.42 \times 10^{24} \text{ atoms}}$$

Thus, the answer is **C**.

2. Because the same number of moles must remain before and after dilution, we can set up the following equation:

$$\begin{aligned} M_1 V_1 &= M_2 V_2 \\ (1.50 \text{ M})(V_1) &= (0.000 \text{ M})(0.500 \text{ L}) \\ (1.50 \text{ M})(V_1) &= 0.150 \text{ M} \cdot \text{L} \\ V_1 &= 0.100 \text{ L} \end{aligned}$$

Thus is **100 mL**, or **A**.

3. We first need to write the reaction in order to perform the stoichiometry. The reaction is as follows:



Now we can do the stoichiometry:

$$10.00 \text{ g Cu} = \frac{1 \text{ mol Cu}}{63.55 \text{ g Cu}} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol Cu}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = \boxed{1.417 \text{ g H}_2\text{O}}$$

Thus, the answer is **B**.

4. Set a standard amount of sample at an arbitrary 100.0 g. Thus, there are 48.41 g Cu, 19.02 g As, and 32.57 g S. We can divide by the molar masses of each of the compounds:

$$\begin{aligned} 48.41 \text{ g Cu} &= \frac{1 \text{ mol Cu}}{63.55 \text{ g Cu}} = 0.7618 \text{ mol} \\ 19.02 \text{ g As} &= \frac{1 \text{ mol As}}{74.92 \text{ g As}} = 0.2539 \text{ mol} \\ 32.57 \text{ g S} &= \frac{1 \text{ mol S}}{32.06 \text{ g S}} = 1.016 \text{ mol} \end{aligned}$$

Dividing by the lowest number, there there will be the ratio of Cu : As : S = 3 : 1 : 4. Thus, the formula is **Cu₃AsS₄**, or **C**.

National Standardized Chemistry Exam

Ying-Ying Zheng



National Standardized Chemistry Exam:

This book delves into National Standardized Chemistry Exam. National Standardized Chemistry Exam is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into National Standardized Chemistry Exam, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to National Standardized Chemistry Exam
- Chapter 2: Essential Elements of National Standardized Chemistry Exam
- Chapter 3: National Standardized Chemistry Exam in Everyday Life
- Chapter 4: National Standardized Chemistry Exam in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of National Standardized Chemistry Exam. This chapter will explore what National Standardized Chemistry Exam is, why National Standardized Chemistry Exam is vital, and how to effectively learn about National Standardized Chemistry Exam.

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6. In chapter 5, the author will draw a conclusion about National Standardized Chemistry Exam. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of National Standardized Chemistry Exam.

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