

NATURAL SCIENCES

Gr. 9 - LESSON PLAN - MATTER & MATERIALS

TOPIC: COMPOUNDS		
Sub-topic: Names of Compounds		Duration: 1 Hour
CONTENT & CONCEPTS (CAPS p63)		
KEY CONCEPTS <ol style="list-style-type: none"> 1. Chemical reaction 2. Chemical formula 3. Names of compounds according to their elements. 		
Specific Aims:	Specific Aim 1: 'Doing Science'	X
	Specific Aim 2: 'Knowing the subject content and making connections'	X
	Specific Aim 3: 'Understanding the uses of Science'	
LESSON OBJECTIVES: <ol style="list-style-type: none"> 1. Learners will be able to write the names of compounds as indicated on page 63 of the CAPS Document. 2. Learners will be able to make models of different compounds 		
RESOURCES REQUIRED: <ul style="list-style-type: none"> • Periodic Table Poster • Siyavula Sasol Inzalo (Learner book) page 163 		
TEACHING & LEARNING ACTIVITY: Introduction <u>Briefly review and revise concepts dealt with in Grade 8, focusing on compounds</u> As each element has a unique name for example: <ul style="list-style-type: none"> - Hydrogen - Lithium Remember the first 20 elements learnt in grade 8. How did you memorize them? Do you remember Mnemonics... He - H; He - He; Lies - Li; Because - Be; Boys - B; Can - C; Not - N; Operate - O; Fireplaces - F; New - Ne; Nation - Na; Might - Mg; Also - Al; Sign - Si; Peace - P; Security - S; Clause - Cl A - Ar; King - K; Can - Ca LEARNERS MAY RECITE THIS MNEMONIC AS A WAY OF MEMORIZING THE FIRST 20 ELEMENTS. <u>The teacher provides the following explanation to the learners</u> Each compound has a unique name. How is a compound formed? <ul style="list-style-type: none"> - A compound is formed when TWO or more different elements combine chemically. - Each atom in a compound are held by chemical bonds - Compounds are formed during chemical reaction (when elements/ compounds combine to form new substances) - Each compound can be written as a formula, for example, H₂O; NaCl; CuSO₄ What is a formula ? <ul style="list-style-type: none"> - This is a notation that shows which elements the compound is made up of. - It indicates the total number of atoms in one unit of the compound - It indicates the ratio in which atoms of different elements are bonded to each other - It gives the total number of atoms in one unit of a compound. For example, the formula for Copper Sulphate is CuSO ₄		

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Table of Contents Natural Science Caps Lesson Plan

1. Understanding the eBook Natural Science Caps Lesson Plan
 - The Rise of Digital Reading Natural Science Caps Lesson Plan
 - Advantages of eBooks Over Traditional Books
2. Identifying Natural Science Caps Lesson Plan
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Natural Science Caps Lesson Plan
 - User-Friendly Interface
4. Exploring eBook Recommendations from Natural Science Caps Lesson Plan
 - Personalized Recommendations
 - Natural Science Caps Lesson Plan User Reviews and Ratings
 - Natural Science Caps Lesson Plan and Bestseller Lists
5. Accessing Natural Science Caps Lesson Plan Free and Paid eBooks
 - Natural Science Caps Lesson Plan Public Domain eBooks
 - Natural Science Caps Lesson Plan eBook Subscription Services
 - Natural Science Caps Lesson Plan Budget-Friendly Options

6. Navigating Natural Science Caps Lesson Plan eBook Formats
 - ePub, PDF, MOBI, and More
 - Natural Science Caps Lesson Plan Compatibility with Devices
 - Natural Science Caps Lesson Plan Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Natural Science Caps Lesson Plan
 - Highlighting and Note-Taking Natural Science Caps Lesson Plan
 - Interactive Elements Natural Science Caps Lesson Plan
8. Staying Engaged with Natural Science Caps Lesson Plan
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Natural Science Caps Lesson Plan
9. Balancing eBooks and Physical Books Natural Science Caps Lesson Plan
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Natural Science Caps Lesson Plan
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Natural Science Caps Lesson Plan
 - Setting Reading Goals Natural Science Caps Lesson Plan
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Natural Science Caps Lesson Plan
 - Fact-Checking eBook Content of Natural Science Caps Lesson Plan
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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