

DEPARTMENT OF EDUCATION
REPUBLIC OF SOUTH AFRICA
NATIONAL CERTIFICATE
ELECTROTECHNICS N5
TIME: 3 HOURS
MARKS: 100

QUESTION 1

1.1 By reversing either the field \checkmark or armature connections. \checkmark (2)

1.2 They are placed in slots in the pole shoes. \checkmark They are connected in series with the armature windings. \checkmark (2)

1.3 $V = 550 \text{ V}$ $P = 44 \text{ kW}$ $N = 1500 \text{ rpm}$ $\eta = 85\% = 0.85$
 $R_a = 0.32$ $I_f = 1.8 \text{ A}$ $I_a R_b = 3.5 \text{ V}$

1.3.1 $\text{Input} = \frac{\text{Output}}{\eta} = 44 / 0.85 = 51.765 \text{ kW} \checkmark$

$I_L = P / V = 51765 / 550 \checkmark = 94.12 \text{ A} \checkmark$ (3)

1.3.2 $\text{Output} = \frac{2 \pi N T}{60}$

$44000 = \frac{2 \pi 1500 T}{60} \checkmark$

$T = \frac{44000 \times 60}{2 \pi 1500} \checkmark$

$= 280.25 \text{ Nm} \checkmark$ (3)

1.3.3 $\text{Starting current} = 1.75 \text{ (full load)} = 1.75 \times 94.12 = 164.7 \text{ A} \checkmark$

$I_a = I - I_f = 164.7 - 1.8 = 162.9 \text{ A} \checkmark$

$V = E + V_b + V_a \text{ (at start)} \quad E = 0 \quad V_a = I_a (R_a + x)$

$550 = 0 + 3.5 + 162.9 (0.32 + x) \checkmark$

$x = \frac{550 - 55.63}{16.9}$

$= 2.035 \text{ ohms} \checkmark$

(4)

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h₂o

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redox equations oxidation state method ion electron method the latter is easier to use with redox reactions in aqueous solution and if necessary can be adapted to many situations that are not in aqueous solution

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-} + 2\text{NO}_3^- + 2\text{MnO}_2 + \text{Cr}_2\text{O}_7^{2-} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method 8 $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow \text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}_2\text{Te}$ 10 $\text{ReO}_4^- + 11\text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H_2O H^+ and e^- in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 + 2\text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 no 2 and $\text{C}_6\text{H}_8\text{O}_6$ b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ h and NO_2 c h and

