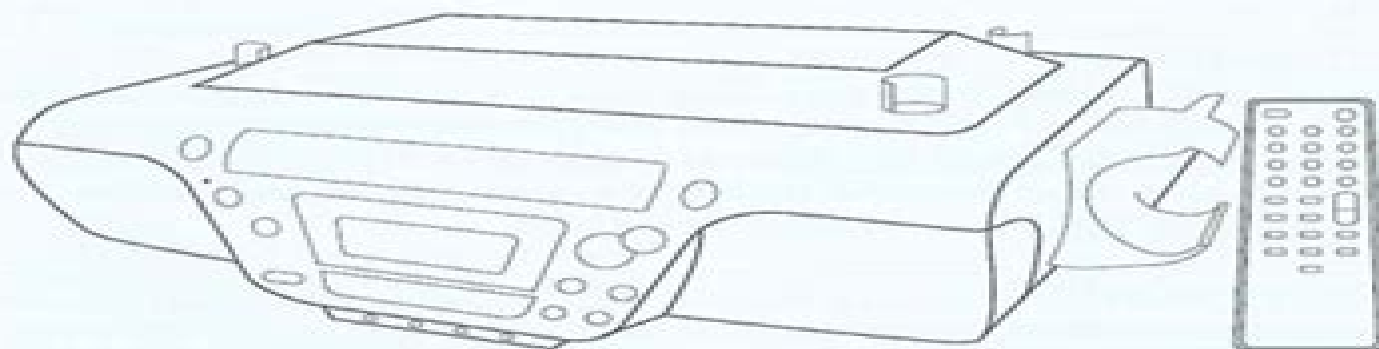


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temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ J/g} \cdot \text{C} \times 110 \text{ g} \times 26.6 \text{ C} = 1210 \text{ J}$
 $550 \text{ J} = \Delta h \cdot q_{\text{water}} = 550 \text{ J}$

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 $14.7 \text{ J} = 3.14 \Delta H_{\text{rxn}} = n \cdot q_{\text{rxn}} = n \cdot q_{\text{calorimeter}} = m \cdot c \cdot \Delta T$

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