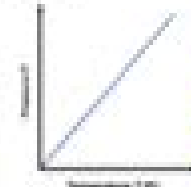
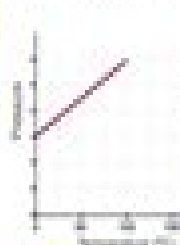
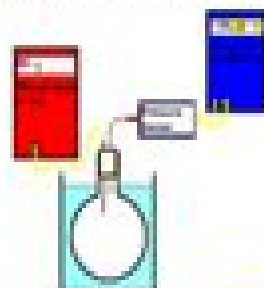


## GAS LAWS SUMMARY

## 1. Pressure and Temperature

(Volume constant, mass constant)



Pressure is directly proportional to Temperature

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$

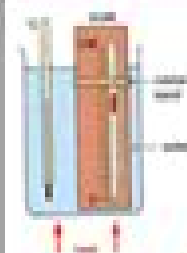
## Extra Notes

Temp ↑ particles have greater speed, and greater KE

Particles collide with container walls more often and with greater force → since P ∝ F/A, Pressure ↑

## 2. Volume and Temperature

(Pressure constant, mass constant)



Volume is directly proportional to Temperature

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

## Extra Notes

Volume ↑ particles collide with container walls more often

Small force increases → since P ∝ F/A, Pressure ↑

## 3. Pressure and Volume

(Temperature constant, mass constant)



$$P_1 V_1 = P_2 V_2$$

## Extra Notes

Volume ↑ particles collide with container walls more often

Small force increases → since P ∝ F/A, Pressure ↑

Pressure is the force per unit area

$$P = \frac{F}{A}$$

p = the pressure (N/m²) F = the force (N) A = the area (m²)

Combined Gas Equation

$$\frac{P}{T} = \text{constant}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

Degrees Celsius to Kelvin: + 273

Kelvin to Degrees Celsius: - 273

Absolute Zero: 0K = 0 °C

 There is no negative on the Kelvin scale  
Temperature is a measure of the mean kinetic energy of particles

SPECIFIC HEAT CAPACITY (c): Heat energy required to raise the temperature of 1 kg of a substance by 1°C

$$E_h = c m \Delta T$$

 E<sub>h</sub>: heat energy (J)

 c: specific heat capacity (J kg<sup>-1</sup> °C<sup>-1</sup>)

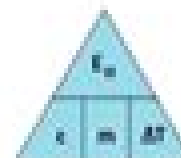
m: mass (kg)

ΔT: change in Temperature (°C)

 Specific heat capacity values are different for all materials  
- values found in data sheet

E = P t can be used to find heat energy supplied by appliances

Heat energy will usually be lost to the surroundings - not all will be transferred to the substance being heated



Low specific heat capacity -

not much energy to heat (heats up quickly for constant power) but loses heat quickly

High specific heat capacity -

lots of energy to heat (heats up slowly for constant power) but loses heat slowly (retains heat better)

Specific Latent Heat (L): Heat energy required to change the state of 1 kg of a substance at the same temperature

$$E_h = m L$$

 E<sub>h</sub>: heat energy (J)

m: mass (kg)

L: specific latent heat of fusion / vaporisation

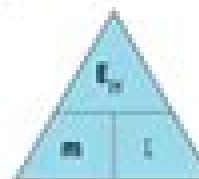
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 Specific Latent heat of fusion, L<sub>f</sub>:

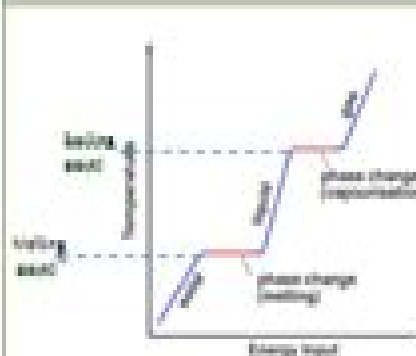
Solid → Liquid

 Specific Latent heat of vaporisation, L<sub>v</sub>:

Liquid → Gas



The same material requires different quantities of heat to change the state of unit mass from solid to liquid (fusion) and to change the state of unit mass from liquid to gas (vaporisation)



During change of state

- temperature remains constant

For change of temperature:

$$E_h = c m \Delta T$$

For change of state:

$$E_h = m L$$

"specific heat" = per kg diagram

# Memorandum Of Engineering Physics N5 For November 2013

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## **Table of Contents Memorandum Of Engineering Physics N5 For November 2013**

1. Understanding the eBook Memorandum Of Engineering Physics N5 For November 2013
  - The Rise of Digital Reading Memorandum Of Engineering Physics N5 For November 2013
  - Advantages of eBooks Over Traditional Books
2. Identifying Memorandum Of Engineering Physics N5 For November 2013
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Memorandum Of Engineering Physics N5 For November 2013
  - User-Friendly Interface
4. Exploring eBook Recommendations from Memorandum Of Engineering Physics N5 For November 2013
  - Personalized Recommendations
  - Memorandum Of Engineering Physics N5 For November 2013 User Reviews and Ratings
  - Memorandum Of Engineering Physics N5 For November 2013 and Bestseller Lists
5. Accessing Memorandum Of Engineering Physics N5 For November 2013 Free and Paid eBooks
  - Memorandum Of Engineering Physics N5 For November 2013 Public Domain eBooks
  - Memorandum Of Engineering Physics N5 For November 2013 eBook Subscription Services

- Memorandum Of Engineering Physics N5 For November 2013 Budget-Friendly Options
- 6. Navigating Memorandum Of Engineering Physics N5 For November 2013 eBook Formats
  - ePub, PDF, MOBI, and More
  - Memorandum Of Engineering Physics N5 For November 2013 Compatibility with Devices
  - Memorandum Of Engineering Physics N5 For November 2013 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Memorandum Of Engineering Physics N5 For November 2013
  - Highlighting and Note-Taking Memorandum Of Engineering Physics N5 For November 2013
  - Interactive Elements Memorandum Of Engineering Physics N5 For November 2013
- 8. Staying Engaged with Memorandum Of Engineering Physics N5 For November 2013
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Memorandum Of Engineering Physics N5 For November 2013
- 9. Balancing eBooks and Physical Books Memorandum Of Engineering Physics N5 For November 2013
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Memorandum Of Engineering Physics N5 For November 2013
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Memorandum Of Engineering Physics N5 For November 2013
  - Setting Reading Goals Memorandum Of Engineering Physics N5 For November 2013
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Memorandum Of Engineering Physics N5 For November 2013
  - Fact-Checking eBook Content of Memorandum Of Engineering Physics N5 For November 2013
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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