

**Question 9. How many oxygen atoms are there in 18 g of water?**

**Answer:** Molar mass of water is 18 g/mol.

Number of oxygen atoms in 18 g of water =  $6.02 \times 10^{23}$

**Question 10. Name two factors that introduce uncertainty into measured figures.**

**Answer:** (i) Reliability of measuring instrument.

(ii) Skill of the person making the measurement.

**Question 11. State Avogadro's law.**

**Answer:** Equal volumes of all gases under the conditions of same temperature and pressure contain the same number of molecules.

**Question 12. How are 0.5 ml of NaOH different from 0.5 M of NaOH?**

**Answer:** 0.5 ml of NaOH means 0.5 mole (20.0 g) of NaOH, 0.5M of NaOH means that 0.5 mole (20.0g) of NaOH are dissolved in 1L of its solution.

**Question 13. What is one a.m.u. or one 'u'?**

**Answer:** 1 a.m.u. or 1 u =  $1/12$  th mass of an atom of carbon 12.

**Question 14. What is the number of significant figures in  $1.050 \times 10^4$ ?**

**Answer:** Four.

## **II. Short Answer Type Questions**

**Question 1. Define molality. How does molality depend on temperature?**

**Answer:** Molality is defined as the moles of solute per kilogram of solvent.

# Mastering Chemistry Homework Answers Chapter 1

**Julia Schneider**



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