

**NOTE:**

- If a candidate answers a question TWICE, only mark the FIRST attempt.
- If a candidate has crossed out an attempt of a question and not redone the question, mark the crossed out version.
- Consistent Accuracy applies in all aspects of the marking memorandum.

**QUESTION 1**

1.1.1	$x(x+1) = 6$ $x^2 + x = 6$ $x^2 + x - 6 = 0$ $(x+3)(x-2) = 0$ $x = -3 \text{ or } 2$ <p><b>OR</b></p> $x^2 + x - 6 = 0$ $x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$ $= \frac{-1 \pm \sqrt{1^2 - 4(1)(-6)}}{2(1)}$ $x = -3 \text{ or } 2$	<div style="border: 1px solid black; padding: 5px; margin-bottom: 10px;"> <b>Note:</b> Answers by inspection: award 3/3 marks         </div> <div style="border: 1px solid black; padding: 5px; margin-bottom: 10px;"> <b>Note:</b> Answer only of <math>x = 2</math> : award 1/3 marks         </div> <div style="border: 1px solid black; padding: 5px;"> <b>Note:</b> If candidate converts equation to linear: award 0/3 marks         </div>	<div style="display: flex; justify-content: space-between;"> <div>             ✓ standard form              ✓ factors              ✓ answers           </div> <div>(3)</div> </div> <div style="display: flex; justify-content: space-between; margin-top: 20px;"> <div>             ✓ standard form              ✓ substitution into correct formula              ✓ answers           </div> <div>(3)</div> </div>
1.1.2	$3x^2 - 4x = 8$ $3x^2 - 4x - 8 = 0$ $x = \frac{-b \pm \sqrt{b^2 - 4ac}}{2a}$ $= \frac{-(-4) \pm \sqrt{(-4)^2 - 4(3)(-8)}}{2(3)}$ $= \frac{4 \pm \sqrt{16 + 96}}{6}$ $= \frac{4 \pm \sqrt{112}}{6}$ $= \frac{2 \pm 2\sqrt{7}}{3}$ $= 2,43 \text{ or } -1,10$	<div style="border: 1px solid black; padding: 5px; margin-bottom: 10px;"> <b>Note:</b> If candidate uses incorrect formula: maximum 1/4 marks (for standard form)         </div> <div style="border: 1px solid black; padding: 5px;"> <b>Note:</b> If an error in subs and gets: <math>\frac{4 \pm \sqrt{-80}}{6}</math> and states "no solution": maximum 3/4 marks  If doesn't conclude with "no solution": maximum 2/4 marks         </div>	<div style="display: flex; justify-content: space-between;"> <div>             ✓ standard form                ✓ substitution into correct formula                ✓ <math>\sqrt{112}</math>              ✓ <math>\frac{4 \pm \sqrt{112}}{6}</math> or              decimal answer           </div> <div>(4)</div> </div>

# Mathematics Paper 1 Feb Mar 2014 Memorandum

**Jiyuan Zhang**



## **Mathematics Paper 1 Feb Mar 2014 Memorandum:**

The Scientific Journal Alex Csiszar, 2018-06-25 Not since the printing press has a media object been as celebrated for its role in the advancement of knowledge as the scientific journal From open communication to peer review the scientific journal has long been central both to the identity of academic scientists and to the public legitimacy of scientific knowledge But that was not always the case At the dawn of the nineteenth century academies and societies dominated elite study of the natural world Journals were a relatively marginal feature of this world and sometimes even an object of outright suspicion The Scientific Journal tells the story of how that changed Alex Csiszar takes readers deep into nineteenth century London and Paris where savants struggled to reshape scientific life in the light of rapidly changing political mores and the growing importance of the press in public life The scientific journal did not arise as a natural solution to the problem of communicating scientific discoveries Rather as Csiszar shows its dominance was a hard won compromise born of political exigencies shifting epistemic values intellectual property debates and the demands of commerce Many of the tensions and problems that plague scholarly publishing today are rooted in these tangled beginnings As we seek to make sense of our own moment of intense experimentation in publishing platforms peer review and information curation Csiszar argues powerfully that a better understanding of the journal's past will be crucial to imagining future forms for the expression and organization of knowledge

An Age of Accountability John L. Rury, 2023-10-13 An Age of Accountability highlights the role of test based accountability as a policy framework in American education from 1970 to 2020 For more than half a century the quest to hold schools and educators accountable for academic achievement has relied almost exclusively on standardized assessment The theory of change embedded in almost all test based accountability programs held that assessment with stipulated consequences could lead to major improvements in schools This was accomplished politically by proclaiming lofty goals of attaining universal proficiency and closing achievement gaps which repeatedly failed to materialize But even after very clear disappointments no other policy framework has emerged to challenge its hegemony The American public today has little confidence in institutions to improve the quality of goods and services they provide especially in the public sector As a consequence many Americans continue to believe that accountability remains a vital necessity even if educators and policy scholars disagree

**Click Here to Kill Everybody** Bruce Schneier, 2018-09-04 Sober lucid and often wise Nature The Internet is powerful but it is not safe As smart devices proliferate the risks will get worse unless we act now From driverless cars to smart thermostats from autonomous stock trading systems to drones equipped with their own behavioral algorithms the Internet now has direct effects on the physical world Forget data theft cutting edge digital attackers can now literally crash your car pacemaker and home security system as well as everyone else's In Click Here to Kill Everybody best selling author Bruce Schneier explores the risks and security implications of our new hyper connected era and lays out common sense policies that will allow us to enjoy the benefits of this omnipotent age without falling prey to the consequences of its

insecurity     Military Cost-Benefit Analysis Francois Melese, Anke Richter, Binyam Solomon, 2015-03-27 This is the first comprehensive book on Military Cost Benefit Analysis and provides novel approaches to structuring cost benefit and affordability analysis amidst an uncertain defense environment and cloudy fiscal prospects Lifting the veil on military Cost Benefit Analysis this volume offers several new practical tools designed to guide defense investments and divestments combined with a selection of real world applications The widespread employment of Cost Benefit Analysis offers a unique opportunity to transform legacy defense forces into efficient effective and accountable 21st century organizations A synthesis of economics statistics and decision theory CBA is currently used in a wide range of defense applications in countries around the world i to shape national security strategy ii to set acquisition policy and iii to inform critical investments in people equipment infrastructure services and supplies As sovereign debt challenges squeeze national budgets and emerging threats disrupt traditional notions of security this volume offers valuable tools to navigate the political landscape meet calls for fiscal accountability and boost the effectiveness of defense investments to help guarantee future peace and stability A valuable resource for scholars practitioners novices and experts this book offers a comprehensive overview of Military Cost Benefit Analysis and will appeal to anyone interested or involved in improving national security and will also be of general interest to those responsible for major government programs projects or policies     **Guide to Microforms in Print** ,1988

*Aero/space Engineering* ,1958     U.S. Government Research & Development Reports ,1971     **Government Reports Announcements & Index** ,1992     The Journal of Biolaw & Business ,2004     **Who's who Among Students in American Universities and Colleges** Henry Pettus Randall,1977     Modern Electronics ,1988

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web oct 22 2019 welding society aws are voluntary consensus standards that have been developed in accordance with the rules of the american national standards institute ansi when aws american national standards are either incorporated in or the a2 committee on definitions and symbols was formed by the american welding society to

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regarding welding symbols and it makes the following distinction weld symbol the symbol representing the type of weld it is placed on the reference line and when used is part of the welding symbol see welding symbol

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web jul 8 2020 about welding symbols and aws a2 4 2020 a welding symbol is the graphical representation of the specifications for producing a welded joint and while the horizontal reference line and arrow are required the symbol can consist of several elements the arrow side is the side of the joint to which the arrow of the welding

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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized  $\text{Ca(OH)}_2$  when calcium is put in water reacts as follows  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant  $K$  is related to  $\Delta G$  for a process through eqn 1 and as long as  $\Delta H$  and  $\Delta S$  are independent of temperature eqn 1 can be combined with eqn 2 to determine  $\Delta H$  and  $\Delta S$  from the temperature dependence of  $K$

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web your primary objective in this experiment is to test a saturated solution of calcium hydroxide and use your observations and measurements to calculate the ksp of the compound you will do this by titrating the prepared  $\text{Ca(OH)}_2$  solution with a standard hydrochloric acid solution

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide  $\text{Ca(OH)}_2$  by measuring the concentration of  $\text{Ca(OH)}_2$  in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of  $\text{Ca(OH)}_2$  is represented in equation form as shown below  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts

paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

*ksp calcium hydroxide chem132 determination of the ksp of calcium* - Jun 24 2022

web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution  $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^-$   $5 \times 10^{-2} \text{ M}$  calcium ions  $5 \times 10^{-2} \text{ M}$  hydroxide ions calculate calcium hydroxide Ksp using the concentrations of calcium ions and hydroxide ions

**solubility of calcium hydroxide wpmu dev** - Oct 29 2022

web madison bowe elana keller april 15 2014 introduction the objective of this lab was to measure the Ksp of a slightly soluble basic salt and how solubility changes with temperature Ksp is another form of equilibrium expression that deals with solubility the reaction that will take place is  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

**solved experiment 8 determining Ksp for calcium hydroxide** - Aug 27 2022

web question experiment 8 determining Ksp for calcium hydroxide post lab report sheet data table 2 points trial 1 trial 2 HCl molarity 0.050 M volume of filtered  $\text{Ca(OH)}_2$  solution mL 5.00 5.00 3.71 13.75 equivalence point volume mL OH<sup>-</sup> in saturated  $\text{Ca(OH)}_2$  M  $\text{Ca}^{2+}$  in saturated  $\text{Ca(OH)}_2$  M Ksp of  $\text{Ca(OH)}_2$  average Ksp of  $\text{Ca(OH)}_2$

chemistry 1210 lab 11 experiment 11 determining the Ksp of calcium - Sep 08 2023

web the chemical equation used was  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  the solubility product was determined to be Ksp  $\text{Ca}^{2+} \text{OH}^-$  from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

18.1 solubility product constant Ksp chemistry libretexts - Feb 01 2023

web jul 12 2023 because the concentration of a pure solid such as  $\text{Ca}_3\text{PO}_4$  is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore  $\text{K}_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]$

**post lab report 8 determining the Ksp of calcium hydroxide** - Jul 06 2023

web lab report 8 determination of calcium hydroxide the purpose of this was to titrate a saturated  $\text{Ca(OH)}_2$  solution with a standard HCl solution and find the Ksp or the OH<sup>-</sup> concentration was found and then the concentration was found and used for this experiment a titration of calcium hydroxide with hydrochloric

*solved determining the Ksp of calciumhydroxide* - Apr 22 2022

web question determining the Ksp of calciumhydroxide  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  for this experiment the titration of calciumhydroxide with hydrochloric acid was investigated determining the Ksp of calcium hydroxide  $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$  for this experiment the titration of calcium



solved chem 182 experiment 8 determining the ksp of calcium chegg - Sep 27 2022

web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of  $\text{Ca(OH)}_2$  is represented in equation form lution of  $\text{Ca(OH)}_2$  on s nronic solid that is sparingly soluble in water a saturated aqueous as shown below

determination of the ksp of calcium hydroxide autosaved - Jun 05 2023

web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for  $\text{Ca(OH)}_2$  was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

exp 4 determining the solubility product constant ksp of calcium - Jul 26 2022

web plasma membrane homework symbiosis practice worksheet benzoin condensation grade a determining the ksp of calcium hydroxide after reading the lab complete items a b c and d title purpose chemicals and equipment and summary of procedure as described on page 10 of exp 1 on an 8 1/2 x 11 sheet of paper

**lab 9 determination of the solubility product ksp chegg com** - May 24 2022

web question lab 9 determination of the solubility product ksp of calcium hydroxide introduction ionic compounds that are classified as insoluble based on solubility rules are actually slightly soluble each of these insoluble compounds actually dissolves to

experimental determination of the solubility product of calcium hydroxide - Nov 29 2022

web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the  $\text{Ca}^{2+}$  and  $\text{OH}^-$  concentrations in solution

**ksp of calcium hydroxide lab report for determining the ksp of** - Oct 09 2023

web calculate the ksp for calcium hydroxide  $\text{ksp Ca(OH)}_2 = 5 \times 10^{-6}$  find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is  $5 \times 10^{-6}$