

## Stoichiometry Worksheet #2 (mole-mass, mass-mole problems)



- a. If 15.0g of  $\text{N}_2\text{O}_4$  was produced, how many moles of  $\text{O}_2$  were required?

$$\frac{15.0\text{g N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} \times \frac{1\text{ mol N}_2\text{O}_4}{1\text{ mol N}_2\text{O}_4} \times \frac{2\text{ mol O}_2}{1\text{ mol N}_2\text{O}_4} = 0.326\text{ mol O}_2$$

- b. If  $4.0 \times 10^{-3}$  moles of oxygen reacted, how many grams of  $\text{N}_2$  were needed?

$$\frac{4.0 \times 10^{-3}\text{ mol O}_2}{2\text{ mol O}_2} \times \frac{1\text{ mol N}_2}{1\text{ mol N}_2} \times \frac{28.0\text{g N}_2}{1\text{ mol N}_2} = 5.6 \times 10^{-2}\text{ g N}_2$$



- a. What is the mass of potassium nitrate that is produced when 2.04 moles of potassium phosphate react?

$$\frac{2.04\text{ mol K}_3\text{PO}_4}{1\text{ mol K}_3\text{PO}_4} \times \frac{3\text{ mol KNO}_3}{1\text{ mol K}_3\text{PO}_4} \times \frac{101.1\text{ g KNO}_3}{1\text{ mol KNO}_3} = 619\text{g KNO}_3$$

- b. If 5.80g of aluminum phosphate are formed, how many moles of aluminum nitrate reacted?

$$\frac{5.80\text{g AlPO}_4}{122.0\text{g AlPO}_4} \times \frac{1\text{ mol AlPO}_4}{1\text{ mol AlPO}_4} \times \frac{1\text{ mol Al}(\text{NO}_3)_3}{1\text{ mol AlPO}_4} = 4.75 \times 10^{-2}\text{ mol Al}(\text{NO}_3)_3$$



- a. If you have 5.50 mol of  $\text{CaC}_2$ , how much  $\text{C}_2\text{H}_2$  do you get?

$$\frac{5.50\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{1\text{ mol C}_2\text{H}_2}{1\text{ mol CaC}_2} \times \frac{26.0\text{g C}_2\text{H}_2}{1\text{ mol C}_2\text{H}_2} = 143\text{g C}_2\text{H}_2$$

- b. How many moles of water are needed when 65.0g of  $\text{CaC}_2$  have reacted?

$$\frac{65.0\text{g CaC}_2}{64.1\text{g CaC}_2} \times \frac{1\text{ mol CaC}_2}{1\text{ mol CaC}_2} \times \frac{2\text{ mol H}_2\text{O}}{1\text{ mol CaC}_2} = 2.03\text{ mol H}_2\text{O}$$

4. In photosynthesis, water reacts with carbon dioxide to give oxygen and glucose ( $\text{C}_6\text{H}_{12}\text{O}_6$ ). Write and balance the chemical equation. How many moles of  $\text{CO}_2$  are required to make 120.0g of glucose?



$$\frac{120.0\text{g C}_6\text{H}_{12}\text{O}_6}{180.0\text{g C}_6\text{H}_{12}\text{O}_6} \times \frac{1\text{ mol C}_6\text{H}_{12}\text{O}_6}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} \times \frac{6\text{ mol CO}_2}{1\text{ mol C}_6\text{H}_{12}\text{O}_6} = 4.000\text{ mol CO}_2$$

# Mixed Stoichiometry Problem Answers

**J Ma**



## **Mixed Stoichiometry Problem Answers:**

The book delves into Mixed Stoichiometry Problem Answers. Mixed Stoichiometry Problem Answers is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Mixed Stoichiometry Problem Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Mixed Stoichiometry Problem Answers
- Chapter 2: Essential Elements of Mixed Stoichiometry Problem Answers
- Chapter 3: Mixed Stoichiometry Problem Answers in Everyday Life
- Chapter 4: Mixed Stoichiometry Problem Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Mixed Stoichiometry Problem Answers. The first chapter will explore what Mixed Stoichiometry Problem Answers is, why Mixed Stoichiometry Problem Answers is vital, and how to effectively learn about Mixed Stoichiometry Problem Answers.

3. In chapter 2, the author will delve into the foundational concepts of Mixed Stoichiometry Problem Answers. The second chapter will elucidate the essential principles that need to be understood to grasp Mixed Stoichiometry Problem Answers in its entirety.

4. In chapter 3, the author will examine the practical applications of Mixed Stoichiometry Problem Answers in daily life. This chapter will showcase real-world examples of how Mixed Stoichiometry Problem Answers can be effectively utilized in everyday scenarios.

5. In chapter 4, this book will scrutinize the relevance of Mixed Stoichiometry Problem Answers in specific contexts. This chapter will explore how Mixed Stoichiometry Problem Answers is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, the author will draw a conclusion about Mixed Stoichiometry Problem Answers. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Mixed Stoichiometry Problem Answers.

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