

## Section 4.5

# Types of Aqueous Solutions and Solubility

## Solubility Rules for Ionic Compounds in Water

TABLE 4.1 Solubility Rules for Ionic Compounds in Water

Compounds Containing the Following Ions Are Generally Soluble	Exceptions
$\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , and $\text{NH}_4^+$	None
$\text{NO}_3^-$ and $\text{C}_2\text{H}_3\text{O}_2^-$	None
$\text{Cl}^-$ , $\text{Br}^-$ , and $\text{I}^-$	When these ions pair with $\text{Ag}^+$ , $\text{Hg}_2^{2+}$ , or $\text{Pb}^{2+}$ , the resulting compounds are insoluble.
$\text{SO}_4^{2-}$	When $\text{SO}_4^{2-}$ pairs with $\text{Sr}^{2+}$ , $\text{Ba}^{2+}$ , $\text{Pb}^{2+}$ , $\text{Ag}^+$ , or $\text{Ca}^{2+}$ , the resulting compound is insoluble.
Compounds Containing the Following Ions Are Generally Insoluble	Exceptions
$\text{OH}^-$ and $\text{S}^{2-}$	When these ions pair with $\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , or $\text{NH}_4^+$ , the resulting compounds are soluble.
	When $\text{S}^{2-}$ pairs with $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , or $\text{Ba}^{2+}$ , the resulting compound is soluble.
	When $\text{OH}^-$ pairs with $\text{Ca}^{2+}$ , $\text{Sr}^{2+}$ , or $\text{Ba}^{2+}$ , the resulting compound is slightly soluble.
$\text{CO}_3^{2-}$ and $\text{PO}_4^{3-}$	When these ions pair with $\text{Li}^+$ , $\text{Na}^+$ , $\text{K}^+$ , or $\text{NH}_4^+$ , the resulting compounds are soluble.

# Model Rules Of Solubility In Aqueous Solutions Pogil

**Atherton Seidell**



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dissolution rate of a drug or other solute in an aqueous medium The book begins with the factors which govern solubility in general and then looks at aqueous solubility in particular including the properties of liquid mixtures and the thermodynamics of solutions formed from mixing two components The bulk of the book is then devoted to techniques for altering solubility and dissolution rate of organic compounds in aqueous media It discusses in detail the most commonly used solubility enhancers buffers cosolvents surfactants and complexants Each chapter is self contained and emphasizes the details for applying the techniques      *Principles of Solution and Solubility* Kōzō Shinoda,1978

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