

# Chemistry - Hall Lab

## Purpose

The purpose of this laboratory is to determine the value of copper produced in this experiment as a percentage yield.

## Procedure

### Step 1

1. Label three pieces of 100 mL beakers.
2. Put 10.0000 g of solid 100% Zn in beaker 1, 10.0000 g of copper (II) chloride in beaker 2. To beaker 3, pour the solution that beaker 2 has up to the 100 mL mark. Add copper (II) chloride to beaker 3 until the beaker is full again.
3. Add about 10 mL distilled water to beaker 1. Stir it with a stirring rod.
4. Allow it to sit for 10 minutes to 15 min.
5. Place the beaker in the copper chloride solution. Observe the reaction, record your observations. Place the beaker back in the place designated for your reaction.

### Step 2

6. Remove the solids. Place a beaker that precipitates copper. Wash the solids with your distilled water solution. Place the beaker in a beaker with water. Allow the separation of the solids. Place the beaker in the drying oven.
7. Remove the solids from the water solution. Place the precipitate solids about 10 mL of distilled water. Try to have two beakers of the solid copper in your water. Allow your beaker to sit in the solution until the water is clear. Stir the copper with 10 mL of 1.5 M HCl. Remove the beaker with distilled water. Wash the beaker with distilled water in the drying oven.
8. When the dry solids, then check the solids.

### Step 3

9. When the beaker is dry copper. Remove the copper to the place designated for your beaker. Wash your beaker with the dry.

## Table

Mass (Zn and Cu)	100.00 g
Mass (Zn and Cu) + copper (II) chloride	100.00 g
Mass (Zn and Cu) + copper (II) chloride	100.00 g

# Modeling Chemistry Nail Lab

**Charles Egolf Clifton**



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