

Chemistry - Hall Lab

Purpose

The purpose of this laboratory is to determine the value of copper produced in this experiment as a percentage yield.

Procedure

Step 1

1. Label three pieces of 100 mL beakers.
2. Put 10.0000 g of solid 100% Zn in beaker 1, 10.0000 g of copper (II) chloride in beaker 2. To beaker 3, pour the solution that beaker 2 has up to the 100 mL mark. Add copper (II) chloride to beaker 3 until the beaker is full again.
3. Add about 100 mL distilled water to beaker 1. Stir it with a stirring rod.
4. Allow it to sit for 10 minutes to 15 min.
5. Place the beaker in the copper chloride solution. Observe the reaction, record your observations. Place the beaker back in the place designated for your reaction.

Step 2

6. Remove the solids. Place a beaker that precipitates copper. Wash the solids with your distilled water solution. Place the beaker in a beaker with water. Allow the separation of the solids. Place the beaker in the drying oven.
7. Remove solution from the 100 mL beaker. Place the precipitate into about 100 mL of distilled water. Try to have two beakers of the solid copper in your water. Pour the solution into a beaker with distilled water, stir the copper with 10 mL of 1.5 M HCl. Remove the beaker with distilled water. Wash the beaker with distilled water in the drying oven.
8. When the dry solids, then check the solids.

Step 3

9. When the beaker is dry copper. Remove the copper to the place designated for your beaker. Wash your beaker with the dry.

Table

Mass (Zn) and Beaker	100.00 g
Mass (Zn) and Beaker + copper (II) chloride	100.00 g
Mass solids Beaker reaction	100.00 g

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M Woodhall



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