

General Chemistry - Unit 1 (Part 1) Review WS - **ANSWERS**

Directions: Complete the following questions on a separate piece of paper.

LT #2: Understand the experiments and the historical theories that lead to the modern atomic theory.

- Describe Dalton's beliefs about the atom and chemical reactions with atoms.
All elements are composed of atoms; all atoms of a given element are identical, that atoms of different elements are different; chemical compounds are formed when atoms combine; atoms are not created or destroyed in chemical reactions (indivisible); solid sphere.
- Describe the atom according to Thomson's theory.
Thomson found that the atom consisted of a positive mass but there were negative pieces scattered throughout the positive area (Plum Pudding Model). He discovered the electron (negatively charged).
- Describe Rutherford's gold foil experiment. How did he change the model of the atom after the experiment?
Rutherford discovered the nucleus contained protons by shooting alpha particles (+) through gold foil. He found that most went through but some bounced back. He found the nucleus, which is very, very small and dense is found in the center of the atom, with the electrons around the outside.
- What did Chadwick change about the atomic model?
Chadwick (along with the Curies) found that the nucleus also contained the neutrons, which have approximately the same mass as the proton, but have no charge.

LT #3: Understand how subatomic particles determine the properties of an atom or ion.

- List the three subatomic particles, their location, their charge and their relative mass.
*Proton—found in nucleus, + charge, 1 amu (atomic mass unit)
Neutron—found in nucleus, no charge, 1 amu
Electron—cloud surrounding the nucleus, - charge, approx. no mass*
- What is a cation and how is it formed?
A cation is a positive ion and is formed when an atom loses one or more electrons.
 - What is an anion and how is it formed?
An anion is a negative ion and is formed when an atom gains one or more electrons.
- Complete the following charts for atoms and ions:

ATOM	ISOTOPE SYMBOL	ATOMIC #	MASS #	#PROTONS	#ELECTRONS	#NEUTRONS
Sodium - 23	²³ ₁₁ Na	11	23	11	11	12
Aluminum - 27	²⁷ ₁₃ Al	13	27	13	13	14
Chlorine - 35	³⁵ ₁₇ Cl	17	35	17	17	18
Barium - 136	¹³⁶ ₅₆ Ba	56	136	56	56	80

SYMBOL	MASS	#PROTONS	#ELECTRONS	#NEUTRONS	NET CHARGE
K ⁺¹	39	19	18	20	+1
Br ⁻¹	80	35	36	45	-1
Fe ⁺³	56	26	23	30	+3
Se ⁻²	79	34	36	45	-2

- Are all atoms of the same element identical? If not, how can they differ?
No; isotopes differ by # of neutrons. (Isotopes have the same # of protons and electrons.)
- Which of lead's four isotopes is most abundant: lead-204, lead-206, lead-207 or lead-208? Explain how you know which isotope is more abundant. *The average atomic mass for lead is 207.2 amu; therefore, the most abundant isotope is lead-207*

LT #4: Apply the mole concept (molar mass, mole equalities, conversions/dimensional analysis).

- Calculate the molar mass of each of the following atoms or compounds:
 - zinc
65.4 g/mol
 - barium phosphate, Ba₃(PO₄)₂
3(137.3) + 2(31.0) + 8(16.0) = 601.9 g/mol

Modeling Chemistry U6 Wsv2 0 Answers

Chao Zhang



Modeling Chemistry U6 Wsv2 0 Answers:

The book delves into Modeling Chemistry U6 Wsv2 0 Answers. Modeling Chemistry U6 Wsv2 0 Answers is an essential topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Modeling Chemistry U6 Wsv2 0 Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Modeling Chemistry U6 Wsv2 0 Answers
- Chapter 2: Essential Elements of Modeling Chemistry U6 Wsv2 0 Answers
- Chapter 3: Modeling Chemistry U6 Wsv2 0 Answers in Everyday Life
- Chapter 4: Modeling Chemistry U6 Wsv2 0 Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Modeling Chemistry U6 Wsv2 0 Answers. This chapter will explore what Modeling Chemistry U6 Wsv2 0 Answers is, why Modeling Chemistry U6 Wsv2 0 Answers is vital, and how to effectively learn about Modeling Chemistry U6 Wsv2 0 Answers.
3. In chapter 2, the author will delve into the foundational concepts of Modeling Chemistry U6 Wsv2 0 Answers. The second chapter will elucidate the essential principles that need to be understood to grasp Modeling Chemistry U6 Wsv2 0 Answers in its entirety.
4. In chapter 3, the author will examine the practical applications of Modeling Chemistry U6 Wsv2 0 Answers in daily life. The third chapter will showcase real-world examples of how Modeling Chemistry U6 Wsv2 0 Answers can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Modeling Chemistry U6 Wsv2 0 Answers in specific contexts. The fourth chapter will explore how Modeling Chemistry U6 Wsv2 0 Answers is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Modeling Chemistry U6 Wsv2 0 Answers. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Modeling Chemistry U6 Wsv2 0 Answers.

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