

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<hr/> 58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<hr/> 84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<hr/> 159.6

Modeling Instruction 20unit Answers

Aimee Kalisz



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1. Understanding the eBook Modeling Instruction 20unit Answers
 - The Rise of Digital Reading Modeling Instruction 20unit Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Instruction 20unit Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Instruction 20unit Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Instruction 20unit Answers

- Personalized Recommendations
- Modeling Instruction 20unit Answers User Reviews and Ratings
- Modeling Instruction 20unit Answers and Bestseller Lists
- 5. Accessing Modeling Instruction 20unit Answers Free and Paid eBooks
 - Modeling Instruction 20unit Answers Public Domain eBooks
 - Modeling Instruction 20unit Answers eBook Subscription Services
 - Modeling Instruction 20unit Answers Budget-Friendly Options
- 6. Navigating Modeling Instruction 20unit Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Modeling Instruction 20unit Answers Compatibility with Devices
 - Modeling Instruction 20unit Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Modeling Instruction 20unit Answers
 - Highlighting and Note-Taking Modeling Instruction 20unit Answers
 - Interactive Elements Modeling Instruction 20unit Answers
- 8. Staying Engaged with Modeling Instruction 20unit Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Instruction 20unit Answers
- 9. Balancing eBooks and Physical Books Modeling Instruction 20unit Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Instruction 20unit Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Instruction 20unit Answers
 - Setting Reading Goals Modeling Instruction 20unit Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Instruction 20unit Answers

- Fact-Checking eBook Content of Modeling Instruction 20unit Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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