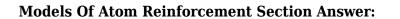
## **Atomic Structure and Chemical Bonds**

## Section 1 Why do atoms combine?

Ä,	The	containing protons	and recetower, in at the	cocotes of an atom and in-	
	surrounded by the		, an area of space around the		
		witheren	travel.		
	16	have a negative charge and do not travel in well defined orbits.			
	2. Each element has a different		100	nisting of a particular	
	mumber of protons, neutrons, and electrons.				
e.	B, Theand		of electrons is	of electrons in the electron cloud-	
	determines the physical and chemical properties of the element.				
	L. Electrons are arranged in different			e different distances from	
	the modeus.				
	2. The an energy level is from the madeus, the more electrons it can hold				
	3. Electrons in the level dissent to the nucleus have		n have the	amount of energy:	
	electrons farthest from the nucleus have the		he	amount of energy.	
Ċ.	Dista from the		can be used?	to undentand energy level	
	L. The is the same as the number of protons (or electrons) in an				
	electrically neutral atom.				
	2. The number of electrons in an element's outremost energy level				
	from left to rig	ht across a period.			
	a. The first energy level is complete with two electrons, and the				
	two efentees	DA.			
	b. The last cle	meet in each of the other_	han .	electrons	
	in its outer energy level.				
	3. Each column in the periodic table contains one element family or group with				
	chemical properties.				
	a. The	in Group	p 8 dio not combine est	sily with other elements	
	Sections their energy levels are stable.				

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