

CHEMICAL EQUILIBRIUM Chapter 9

- (a) Products (b) Reversible reaction
(c) Reactants (d) Irreversible reaction
- Name the reactants in this equation $2\text{H}_2 + \text{O}_2 \xrightarrow[\text{U.T.}]{\text{heat, Pt}} 2\text{H}_2\text{O}$
- (a) H_2O (b) H_2 and O_2 (c) O_2 (d) None of these

QUESTIONS

- The reaction in which the products recombine to form reactants is known as:
- (a) Reversible reactions (b) Reactants
(c) Products (d) Irreversible reactions
- The reactions in which the products recombine to form reactants are called:
- (a) Reactants (b) Reversible reactions
(c) Irreversible reaction (d) Both a and b

In chemical reactions the substances that combine are called:

- Irreversible reaction goes on:
- (a) Never complete (b) Both a and c (c) completion (d) None of these
- Reversible reaction:
- (a) can complete (b) never complete (c) Both a and b (d) None of these
7. The colour of hydrogen iodide (HI) is:
- (a) Blue (b) Grey (c) Purple (d) Colourless
8. The colour of iodine (I_2) is:
- (a) Purple (b) Green (c) Yellow (d) None of these
9. At equilibrium the rate of forward and reverse reaction becomes:
- (a) Greater (b) Normal (c) Equal (d) Lesser
10. Dynamic means reaction is:
- (a) Stops (b) Still continuing (c) Opposite direction (d) Both a and b
11. When reaction ceases to proceed it is called:
- (a) Dynamic equilibrium (b) Static equilibrium
(c) chemical equilibrium (d) none of these
12. Reaction in which reactants react to form products called:
- (a) forward reactions (b) reverse reactions (c) reversible reaction (d) backward reactions
- At initial stage the rate of forward reaction is:
- (a) low (b) very low (c) very fast (d) all of these
14. There are _____ possibilities at equilibrium state.
- (a) three (b) five (c) four (d) two
15. Reverse reaction takes place from:
- (a) left to right (b) right to left (c) both a and b (d) none of these
16. Reverse reactions _____ gradually.
- (a) speeds up (b) negligible (c) slows down (d) none of above
17. Forward reaction takes place from:
- (a) left to right (b) right to left (c) both a and b (d) all of these

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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study with quizlet and memorize flashcards containing terms like reacting sulfuric and nitric acid together produces the nitronium ion alone nitric acid is not a strong enough electrophile to react with bromobenzene sulfuric acid is added to the nitric acid the reverse rxn can overheat and boil easily includes a claisen adaptor and

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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