

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Pd: \_\_\_\_\_

### **MIXED** Stoichiometry Practice Review Problems

STEP 1: The following reaction will be used for this entire worksheet, please complete the balancing before going on.



STEP 2: Calculator the molar mass of each compound in the balanced reaction and write the values above the balanced reaction.

STEP 3: For aluminum sulfate, find the % of the metal in the compound. Please be sure to show your work. (Hint: WS.1 work)

**Al is the only metal, S and O are nonmetals. There are no metalloids in the formula.**

$$\% \text{ Al} = \frac{\text{total mass Al}}{\text{molar mass Al}_2(\text{SO}_4)_3} \times 100 \% = \frac{53.96 \text{ g}}{342.17 \text{ g}} \times 100 \% = 15.77 \%$$

STEP 4: Complete the following problems using **a factor chart with units & chemicals**. These problems could require ONE ratio, TWO ratios or even THREE ratios. Look for clues & use your mole map if necessary.

#### **PROBLEMS**

1) How many **grams of calcium sulfate** are produced from reacting **2.59 moles of aluminum sulfate**?

$$2.59 \text{ mol Al}_2(\text{SO}_4)_3 \times \frac{3 \text{ mol CaSO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{136.15 \text{ g CaSO}_4}{1 \text{ mol CaSO}_4} = 1057.89 \text{ g CaSO}_4$$

moles A  $\longrightarrow$  moles B  $\longrightarrow$  grams B

# Mixed Problems Stoichiometry With Answers

**James E. Brady**



## **Mixed Problems Stoichiometry With Answers:**

**Chemistry** Carson-Dellosa Publishing, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be aligned to current science standards

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**Chemistry Homework** Frank Schaffer Publications, Joan DiStasio, 1996-03 Includes the periodic table writing formulas balancing equations stoichiometry problems and more

**Freshman Chemistry Problems and how to Solve Them** Michell J. Sienko, 1964

**The Practice of Chemistry** Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POC/PowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POC/PowerPoint_Final.ppt) 832KB

**Instructor's manual and solutions to problems to accompany General chemistry: principles and structures** James E. Brady, 1975-04

**Problem Exercises for General Chemistry** George Gilbert Long, Forrest C. Hentz, 1982

**Chemists' Guide to Effective Teaching** Norbert J. Pienta, Melanie M. Cooper, Thomas J. Greenbowe, 2005 Part of the Prentice Hall Series in Educational Innovation for Chemistry this unique book is a collection of information examples and references on learning theory teaching methods and pedagogical issues related to teaching chemistry to college students In the last several years there has been considerable activity and research in chemical education and the materials in this book integrate the latest developments in

chemistry Each chapter is written by a chemist who has some expertise in the specific technique discussed has done some research on the technique and has applied the technique in a chemistry course      *Problem Exercises for General Chemistry* G.Gilbert Long,Forrest C. Hentz,1986-08-14 A question answer supplement to accompany a 1 year general chemistry course This text is a multiple choice format and can be used with any standard general chemistry text The exercises emphasize the importance of problem solving and equation writing in the style used in general chemistry examinations and professional school aptitude examinations such as the MCAT and VCAT      *General Chemistry* James E. Brady,1986      **Chemistry** James E. Brady,Fred Senese,2004-02-04 Publisher Description      **Chemistry** Steven S. Zumdahl,1986      **General Chemistry, Study Guide** James E. Brady,1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems      *Scientific and Technical Aerospace Reports* ,1991      **An Introductory Course of Quantitative Chemical Analysis** Henry Paul Talbot,1908      Modern Chemistry Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001      **Merrill Chemistry** Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998      **Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry** Iris Stovall,Donald J. DeCoste,2000      *Holt Chemistry* Ralph Thomas Myers,2004      **Chemical Principles** Steven S. Zumdahl,1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

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