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### **MIXED** Stoichiometry Practice Review Problems

STEP 1: The following reaction will be used for this entire worksheet, please complete the balancing before going on.



STEP 2: Calculator the molar mass of each compound in the balanced reaction and write the values above the balanced reaction.

STEP 3: For aluminum sulfate, find the % of the metal in the compound. Please be sure to show your work. (Hint: WS.1 work)

**Al is the only metal, S and O are nonmetals. There are no metalloids in the formula.**

$$\% \text{ Al} = \frac{\text{total mass Al}}{\text{molar mass Al}_2(\text{SO}_4)_3} \times 100 \% = \frac{53.96 \text{ g}}{342.17 \text{ g}} \times 100 \% = 15.77 \%$$

STEP 4: Complete the following problems using **a factor chart with units & chemicals**. These problems could require ONE ratio, TWO ratios or even THREE ratios. Look for clues & use your mole map if necessary.

#### **PROBLEMS**

1) How many **grams of calcium sulfate** are produced from reacting **2.59 moles of aluminum sulfate**?

$$2.59 \text{ mol Al}_2(\text{SO}_4)_3 \times \frac{3 \text{ mol CaSO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{136.15 \text{ g CaSO}_4}{1 \text{ mol CaSO}_4} = 1057.89 \text{ g CaSO}_4$$

moles A  $\longrightarrow$  moles B  $\longrightarrow$  grams B

# Mixed Stoichiometry Answers

**Diana Harris**



## Mixed Stoichiometry Answers:

*Computer-assisted Instruction in Chemistry: General approach* Harry B. Mark, Hubert C. MacDonald, 1974

**Instructor's manual and solutions to problems to accompany General chemistry: principles and structures**

James E. Brady, 1975-04      *Molecular Metals* William Hatfield, 2013-03-09 During the past few years there has been intense research activity in the design synthesis and characterization of materials which are formed from molecular precursors and which have high or metal like electrical conductivities i e dcr dT      **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

*The Practice of Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03

Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt) 832KB      *Scientific and Technical Aerospace Reports*, 1991

**Computer-assisted Instruction in Chemistry: General approach** James S. Mattson, Harry B. Mark, Hubert C. MacDonald, 1974      *Proceedings of NECC 1979, National Educational Computing Conference* Diana Harris, 1979

**Chemistry** Jerry A. Bell, American Chemical Society, 2005 The American Chemical Society has launched an activities based student centered approach to the general chemistry course a textbook covering all the traditional general chemistry topics but arranged in a molecular context appropriate for biology environmental and engineering students Written by industry chemists and educators Chemistry combines cooperative learning strategies and active learning techniques with a powerful media supplements package to create an effective introductory text Online description      **Chemistry**, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry The 100 Series science books span grades 5 to 12 The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science The books include engaging grade appropriate activities and clear thumbnail answer keys Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics The series will be aligned to current science standards      **CUET PG Nanoscience**

**(Question Bank 2400 MCQs Chapterwise)** Mocktime Publication, CUET PG Nanoscience Question Bank 2400 MCQs Chapterwise      **Quantitative Chemical Analysis** Daniel C. Harris, Chuck Lucy, 2015-05-29 The gold standard in analytical chemistry Dan Harris Quantitative Chemical Analysis provides a sound physical understanding of the principles of analytical

chemistry and their applications in the disciplines      *Chemistry* Steven S. Zumdahl,1986      **Chemistry Homework** Frank Schaffer Publications,Joan DiStasio,1996-03 Includes the periodic table writing formulas balancing equations stoichiometry problems and more      **Gels** ,1995      **Modern Chemistry** Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2001

*Chemical Principles* Steven S. Zumdahl,1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples      *Chemical Process* Robin Smith,2005-06-10 This book deals with the design and integration of chemical processes emphasizing the conceptual issues that are fundamental to the creation of the process Chemical process design requires the selection of a series of processing steps and their integration to form a complete manufacturing system The text emphasizes both the design and selection of the steps as individual operations and their integration Also the process will normally operate as part of an integrated manufacturing site consisting of a number of processes serviced by a common utility system The design of utility systems has been dealt with in the text so that the interactions between processes and the utility system and interactions between different processes through the utility system can be exploited to maximize the performance of the site as a whole Chemical processing should form part of a sustainable industrial activity For chemical processing this means that processes should use raw materials as efficiently as is economic and practicable both to prevent the production of waste that can be environmentally harmful and to preserve the reserves of raw materials as much as possible Processes should use as little energy as economic and practicable both to prevent the build up of carbon dioxide in the atmosphere from burning fossil fuels and to preserve reserves of fossil fuels Water must also be consumed in sustainable quantities that do not cause deterioration in the quality of the water source and the long term quantity of the reserves Aqueous and atmospheric emissions must not be environmentally harmful and solid waste to landfill must be avoided Finally all aspects of chemical processing must feature good health and safety practice It is important for the designer to understand the limitations of the methods used in chemical process design The best way to understand the limitations is to understand the derivations of the equations used and the assumptions on which the equations are based Where practical the derivation of the design equations has been included in the text The book is intended to provide a practical guide to chemical process design and integration for undergraduate and postgraduate students of chemical engineering practicing process designers and chemical engineers and applied chemists working in process development Examples have been included throughout the text Most of these examples do not require specialist software and can be performed on spreadsheet software Finally a number of exercises have been added at the end of each chapter to allow the reader to practice the calculation procedures      **An Introductory Course of Quantitative Chemical Analysis** Henry Paul Talbot,1897      **Theoretical and Technological Aspects of Crystal Growth** R. Fornari,C. Paorici,1998-03-30 Proceedings of the 10th International Summer School on Crystal Growth Rimini Italy June 1998

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