Name: _

Date:

STOICHIOMETRY PRACTICE WORKSHEET

1. Using the following equation:

Calculate how many grams of iodine are needed to prepare 28.6 grams of ICI by this reaction.

2. Using the following equation:

$$5 \text{ KNO}_2 + 2 \text{ KMnO}_4 + 3 \text{ H}_2 \text{SO}_4 \rightarrow 5 \text{ KNO}_3 + 2 \text{ MnSO}_4 + \text{ K}_2 \text{SO}_4 + 3 \text{ H}_2 \text{O}_4$$

How many moles and grams of $\rm KMnO_4$ are needed for this reaction on 11.4 grams of $\rm KNO_2?$

3. Using the following equation:

How many moles and grams of oxygen (O_2) are needed to react with 56.8 grams of ammonia by this reaction?

4. Using the following equation:

$$NaIO_3 + 6 HI \rightarrow 3 I_2 + NaI + 3 H_2O$$

Calculate the number of moles and the number of grams of iodine (l_2) that can be made this way from 16.4 grams of NaIO₃.

Mixed Stoichiometry Practice

N. S. Nandagopal

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