# Unit 7 - Equilibrium (7-9%)

Some reactions can occur in both forward and reverse directions, sometimes proceeding in each direction simultaneously.

- 7.5 Integ to Equilibrium: Explain the relationship between the occurrence of a reversible chemical or physical process, and the establishment of equilibrium, to experimental observations.
  - Many observable processes are reversible. Examples include evaporation and condensation of water. absorption and description of a gas, or dissolution and precipitation of a salt. Some important reversible chemical processes include the transfer of protons in acid-base reactions and the transfer of electrons in redox reactions.
  - When equilibrium is reached, no observable changes occur in the system. Reactants and products are simultaneously present, and the concentrations or partial pressures of all species remain constant.
  - The equilibrium state is dynamic. The forward and reverse processes continue to occur at equal rates. resulting in no net observable change.
  - Graphs of concentration, partial pressure, or rate of reaction versus time for simple chemical reactions can be used to understand the establishment of chemical equilibrium.

# **Objectives** Equilibrium

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# **Modeling Chemistry Unit 7 Objectives**

**Katrin Zwirglmaier** 

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