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## Chemistry – Unit 7 Worksheet 3

Write balanced chemical equations for the following reactions.

Note: All coefficients of 1 are omitted; in the first 10 items, the coefficients are all 1.

- Ammonia (NH<sub>3</sub>) reacts with hydrogen chloride to form ammonium chloride.
   NH<sub>2</sub> + HCl → NH<sub>4</sub>Cl
- Calcium carbonate decomposes upon heating to form calcium oxide and carbon dioxide.
   CaCO₂ → CaO + CO₂
- Barium oxide reacts with water to form barium hydroxide.
  BaO + H<sub>2</sub>O → Ba(OH)<sub>2</sub>
- Acetaldehyde (CH3CHO) decomposes to form methane (CH<sub>4</sub>) and carbon monoxide.
   CH<sub>2</sub>CHO → CH<sub>4</sub> + CO
- Zinc reacts with copper (II) nitrate to form zinc nitrate and copper.
  Zn + Cu(NO<sub>3</sub>)<sub>2</sub> → Zn(NO<sub>3</sub>)<sub>2</sub> + Cu
- Calcium sulfite decomposes when heated to form calcium oxide and sulfur dioxide.
   CaSO<sub>3</sub> → CaO + SO<sub>2</sub>
- Iron reacts with sulfuric acid (H<sub>2</sub>SO<sub>4</sub>) to form iron (II) sulfate and hydrogen gas.
   Fe + H<sub>2</sub>SO<sub>4</sub> → FeSO<sub>4</sub> + H<sub>2</sub>
- A nitrogen containing carbon compound, C<sub>2</sub>H<sub>6</sub>N<sub>2</sub>, decomposes to form ethane, C<sub>2</sub>H<sub>6</sub>, and nitrogen gas.
   C<sub>2</sub>H<sub>6</sub>N<sub>2</sub> → C<sub>2</sub>H<sub>6</sub> + N<sub>2</sub>
- Phosgene, COCl<sub>2</sub>, is formed when carbon monoxide reacts with chlorine gas.
  CO + Cl<sub>2</sub> → COCl<sub>2</sub>

# **Modeling Chemistry Unit 7 Ws 4**

**Honghua Tan** 

Modeling Chemistry Unit 7 Ws 4:

Whispering the Techniques of Language: An Mental Journey through Modeling Chemistry Unit 7 Ws 4

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