

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate, KClO_3 (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	$2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$			1.2 moles KClO_3 needed
Before:	1.2	0	0	
Change:	-1.2	+1.2	+1.8	
After:	0	1.2	1.8	

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	$\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$				0.200 moles HCl needed
Before:	0.100	?	0	0	
Change:	-0.100	-0.200 +	0.100 +	0.100	
After:	0	?	0.100	0.100	

3. Ethane, C_2H_6 reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	$2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$			
Before:	2.20	xs	0	0
Change:	-2.20	-7.70	+4.40	+6.60
After:	0	xs	4.40	6.60

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

Equation:	$\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$			
Before:	0.0040	XS	0	0
Change:	-0.0040	-0.0080	+0.0040	+0.0080
After:	0	XS	0.0040	0.0080

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

Modeling Chemistry Unit 8 3

**Carnegie Institute of Technology.
College of Fine Arts**



Modeling Chemistry Unit 8 3:

This book delves into Modeling Chemistry Unit 8 3. Modeling Chemistry Unit 8 3 is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Modeling Chemistry Unit 8 3, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Modeling Chemistry Unit 8 3
- Chapter 2: Essential Elements of Modeling Chemistry Unit 8 3
- Chapter 3: Modeling Chemistry Unit 8 3 in Everyday Life
- Chapter 4: Modeling Chemistry Unit 8 3 in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Modeling Chemistry Unit 8 3. The first chapter will explore what Modeling Chemistry Unit 8 3 is, why Modeling Chemistry Unit 8 3 is vital, and how to effectively learn about Modeling Chemistry Unit 8 3.
3. In chapter 2, the author will delve into the foundational concepts of Modeling Chemistry Unit 8 3. This chapter will elucidate the essential principles that need to be understood to grasp Modeling Chemistry Unit 8 3 in its entirety.
4. In chapter 3, this book will examine the practical applications of Modeling Chemistry Unit 8 3 in daily life. The third chapter will showcase real-world examples of how Modeling Chemistry Unit 8 3 can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Modeling Chemistry Unit 8 3 in specific contexts. This chapter will explore how Modeling Chemistry Unit 8 3 is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Modeling Chemistry Unit 8 3. This chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Modeling Chemistry Unit 8 3.

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