

CHE1APL – Environmental Chemistry Workshop 1

This workshop sheet consists of two sections. Section 1 includes questions for which your tutor will be available to check your workings and help you during the workshop session. To aid your learning, make sure that you attempt the workshop questions before coming to class. Section 2 includes short-answer or multiple-choice questions that you can work through in your own time; answers to which will be provided during the workshop.

Section 1.

In this section, you will use some basic mathematics skills that are essential in science to manipulate logarithms and exponential functions, as well as solve simple equations using basic calculus. You can refresh your maths skills by examining the content of Appendix I on "Mathematics for Chemistry: Exponential Notation, Logarithms, Graphing, and Quadratic Equations" in the prescribed textbook. Additional on-line maths help is available from OWL exercises via the LMS for CHE1APL.

Factors Affecting Solubility: Pressure and Temperature (Section 12.4)

1. The Henry's law constant for N_2 in water at 37°C is $6.2 \times 10^{-6} \text{ mol L}^{-1} \text{ kPa}^{-1}$. Calculate the solubility (in mol L^{-1}) of N_2 in water, when the partial pressure of N_2 is 80.7 kPa?

- A. $1.4 \times 10^{-9} \text{ mol L}^{-1}$ B. $6.5 \times 10^{-7} \text{ mol L}^{-1}$ C. $5.0 \times 10^{-4} \text{ mol L}^{-1}$ D. $2.0 \times 10^{-3} \text{ mol L}^{-1}$

Henry's Law: $s = k_H \times p$

solubility (N_2) = Henry's constant $\times$ partial pressure (N_2)

$$s = 6.2 \times 10^{-6} \text{ mol L}^{-1} \text{ kPa}^{-1} \times 80.7 \text{ kPa}$$

$$s = 5 \times 10^{-4} \text{ mol L}^{-1}$$

Colligative Properties: Osmosis (Section 12.6)

2. A solution is prepared by dissolving 4.78 g of an unknown non-electrolyte in enough water to make 375 mL of solution. The osmotic pressure of the solution is 135 kPa at 27°C . Determine the molar mass of the solute? ($R = 8.314 \text{ L kPa K}^{-1} \text{ mol}^{-1}$)

- A. 21.2 g mol^{-1} B. 49.4 g mol^{-1} C. 96.8 g mol^{-1} D. 236 g mol^{-1}

Use osmotic pressure equation.

$$\Pi = icRT$$

Where for a non-electrolyte (such as sucrose) $i=1$

$$135 \text{ kPa} = 1.00 \times c \text{ mol L}^{-1} \times 8.314 \text{ L kPa K}^{-1} \text{ mol}^{-1} \times 300 \text{ K}$$

$$c = (135 \text{ kPa}) / (8.314 \text{ L kPa K}^{-1} \text{ mol}^{-1} \times 300 \text{ K}) = 0.05413 \text{ mol L}^{-1}$$

$$\text{now, } c = (\text{mass/MM}) / V,$$

$$\text{therefore, MM} = \text{mass} / (c \cdot V) = 4.78 \text{ g} / (0.05413 \times 0.375) = 235.5 \text{ g mol}^{-1} = 236 \text{ g mol}^{-1}$$

3. What is the osmotic pressure of a 0.126 mol L^{-1} solution of sodium sulphate (Na_2SO_4) at 25°C ? Assume the van't Hoff factor is 2.51.

- A. 783 kPa B. 65.7 kPa C. 312 kPa D. 936 kPa

Use osmotic pressure equation for an electrolyte

$$\Pi = icRT$$

$$\Pi = 2.51 \times 0.126 \text{ mol L}^{-1} \times 8.314 \text{ L kPa K}^{-1} \text{ mol}^{-1} \times 298 \text{ K}$$

$$\Pi = 784 \text{ kPa}$$

Concentration-Time Relationships: Integrated Rate Equations (Section 18.5)

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