

Kinetics, Equilibrium and Thermodynamics: Practice Test

1. Multiple Choice

1. The primary rate equation is given by: $k[A]^m[B]^n$ where:

- gradient of the graph is $k[A]^m[B]^n$
- gradient of the graph is $k[A]^m[B]^n$
- gradient represents all orders that contribute to the overall order of reaction
- order of the reaction is the sum of the orders of the reactants
- does not support

2. A study of reaction order with:

- by determining the half-life of reaction
- by plotting the rate of reaction
- by determining the activation energy of the reaction
- by plotting concentration against time
- order of the reaction is the sum of the orders of the reactants

3. Equilibrium following reaction:



At equilibrium, $[\text{N}_2\text{O}_4] = 0.04 \text{ mol dm}^{-3}$ and $[\text{NO}_2] = 0.08 \text{ mol dm}^{-3}$

The concentration of N_2O_4 at equilibrium is:

- 0.04
- 0.08
- 0.16
- 0.20
- 0.24

4. A reaction following reaction:



At equilibrium, $[\text{N}_2\text{O}_4] = 0.04 \text{ mol dm}^{-3}$ and $[\text{NO}_2] = 0.08 \text{ mol dm}^{-3}$

The concentration of N_2O_4 at equilibrium is:

- 0.04
- 0.08
- 0.16
- 0.20
- 0.24

5. The reaction mixture is at equilibrium:

- the reaction is at equilibrium
- the reaction is at equilibrium at the moment of observation
- the reaction is at equilibrium
- the reaction is at equilibrium
- the reaction is at equilibrium



6. The reaction mixture is at equilibrium:



7. The reaction mixture is at equilibrium:

- The temperature is high enough to provide enough energy to overcome the activation energy of the reaction
- The temperature is high enough that the reaction mixture is at equilibrium
- The reaction is at equilibrium
- The reaction is at equilibrium

Kinetics And Equilibrium Practice Test

Martha Mackin



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