

# Kinetics of Sodium Bicarbonate Decomposition

A thermal gravimetric method was used to measure rates of decomposition of  $\text{NaHCO}_3$  particles. Such decomposition produces a highly porous  $\text{Na}_2\text{CO}_3$  that reacts with  $\text{SO}_2$  rapidly and completely at moderate temperatures. Hence,  $\text{NaHCO}_3$  decomposition provides a reactant with attractive features for  $\text{SO}_2$  removal. The rapid rate of decomposition combined with the high heat effect prevented determining intrinsic rates by constant temperature runs when the temperature level was above 400 K. However, rising-temperature runs, which allowed time for heat transfer to equilibrate temperatures of the thermocouple and particles, gave reliable results at high temperatures. The activation energy was 102 kJ/mol. Porosimeter data verified the large increase in pore volume (from  $0.03$  to  $0.39 \times 10^{-3} \text{ m}^3/\text{kg}$ ) on converting the  $\text{NaHCO}_3$  particles to  $\text{Na}_2\text{CO}_3$ . First-order kinetics were observed up to high conversions, after which the apparent order decreased. However, the sodium bicarbonate could be completely converted to  $\text{Na}_2\text{CO}_3$ .

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## SCOPE

It has been well established that thermal decomposition of  $\text{NaHCO}_3$  particles can produce a highly porous  $\text{Na}_2\text{CO}_3$  product that reacts rapidly with  $\text{SO}_2$  in waste gases (Bares et al., 1970; Marecek et al., 1970). Also, rates of decomposition and of reaction with  $\text{SO}_2$  have been measured separately on different samples (Subramanian et al., 1972; Gearini et al., 1973; Bares et al., 1970; Marecek et al., 1970).

Our objective was to determine the rate of decompo-

sition of a sample of  $\text{NaHCO}_3$  particles at temperatures from 373 to 473 K over the complete range of conversion. A TGA apparatus was employed since this allowed very small sample masses to be used (leading to accurate rates) and the effect of heat and mass of transfer processes could usually be eliminated. Pore-volume data also were obtained on  $\text{NaHCO}_3$  and  $\text{Na}_2\text{CO}_3$  particles.

## CONCLUSIONS AND SIGNIFICANCE

The rapid rate of decomposition of  $\text{NaHCO}_3$  combined with a large, endothermic heat effect required careful choice of operating conditions to obtain rates unaffected by heat transfer. At temperatures above 400 K this required the TGA experiments to be conducted by the rising-temperature method. The results showed that intrinsic kinetics were first-order in mass of  $\text{NaHCO}_3$  up to high conversions. The retardation of the rate beyond this point may be due to changes in the

nature of the remaining  $\text{Na}_2\text{CO}_3$ , or the structure of the  $\text{Na}_2\text{CO}_3$  product surrounding the unreacted  $\text{Na}_2\text{HCO}_3$ . However, the rate remained high at high conversions so that the  $\text{NaHCO}_3$  was completely converted to the carbonate. The data also showed that the rate decreased when the decomposition temperature was reduced from 473 to 373 K. Corresponding pore-volume data indicated that the surface area was greater for particles decomposed at 373 K.

# Kinetics Of Sodium Bicarbonate Decomposition

**Siddappa N.Byrareddy**



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