

Name Solutions

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Ionic Bonding Worksheet

For each pair of elements below draw an atomic diagram showing electrons in different energy levels. Draw arrows to show where the outer electrons will go during a chemical reaction, then draw the resulting compound. Finally, fill in the table below each reaction. Refer to the sample shown.



| Atoms | Valence electrons | Electron transfer from/to each atom | Ions formed in the product |
|-------|-------------------|-------------------------------------|----------------------------|
| Li    | 1 ve              | transfers $1e^-$ to F               | $\text{Li}^{+}$            |
| F     | 7 ve              | accepts $1e^-$ from Li              | $\text{F}^{-}$             |

  

| Reactions                                                             | Atoms | Valence electrons | Electron transfer from/to each atom | Ions formed in the product |
|-----------------------------------------------------------------------|-------|-------------------|-------------------------------------|----------------------------|
| 1) $\text{Li} + \text{Cl} \Rightarrow \text{LiCl}$                    | Li    | 1 ve              | donates $1e^-$                      | $\text{Li}^{+1}$           |
| $\text{Li} + \text{Cl} \rightarrow [\text{Li}]^{+1} [\text{Cl}]^{-1}$ | Cl    | 1 ve              | accepts $1e^-$                      | $\text{Cl}^{-1}$           |
| 2) $\text{Ca} + \text{O} \Rightarrow \text{CaO}$                      | Ca    | 2 ve              | donates $2e^-$                      | $\text{Ca}^{+2}$           |
| $\text{Ca} + \text{O} \rightarrow [\text{Ca}]^{+2} [\text{O}]^{-2}$   | O     | 6 ve              | accepts $2e^-$                      | $\text{O}^{-2}$            |
| 3) $\text{Be} + \text{F} \Rightarrow \text{BeF}_2$                    | Be    | 2 ve              | donates $2e^-$ to each F            | $\text{Be}^{+2}$           |
| $\text{Be} + \text{F} \rightarrow [\text{Be}]^{+2} [\text{F}]^{-1}$   | F     | 7 ve              | each F accepts $1e^-$ each          | $\text{F}^{-1}$            |
| 4) $\text{Mg} + \text{S} \Rightarrow \text{MgS}$                      | Mg    | 2 ve              | donates $2e^-$ to S                 | $\text{Mg}^{+2}$           |
| $\text{Mg} + \text{S} \rightarrow [\text{Mg}]^{+2} [\text{S}]^{-2}$   | S     | 6 ve              | accepts $2e^-$                      | $\text{S}^{-2}$            |
| 5) $\text{K} + \text{F} \Rightarrow \text{KF}$                        | K     | 1 ve              | donates $1e^-$                      | $\text{K}^{+1}$            |
| $\text{K} + \text{F} \rightarrow [\text{K}]^{+1} [\text{F}]^{-1}$     | F     | 7 ve              | accepts $1e^-$                      | $\text{F}^{-1}$            |

# Ionic And Metallic Bonding Worksheet Answers

**SA Dillow**



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