



CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules by sharing or transferring electrons. This process is called bonding. The resulting molecule is more stable than the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by the delocalization of valence electrons in a metal lattice. Example: Fe , Cu , Al .

Properties of Bonds:

- Covalent bonds are strong and directional.
- Ionic bonds are strong and non-directional.
- Metallic bonds are strong and non-directional.

Science Real

A simplified periodic table of elements, color-coded by groups.

IONIC BONDING

Atoms combine to form ionic compounds by transferring electrons. This process is called ionic bonding. The resulting compound is more stable than the individual atoms.

Types of Ionic Bonds:

- Single Bond:** Formed by the transfer of one electron from one atom to another.
- Double Bond:** Formed by the transfer of two electrons from one atom to another.
- Triple Bond:** Formed by the transfer of three electrons from one atom to another.

Properties of Ionic Bonds:

- Ionic bonds are strong and non-directional.
- Ionic compounds are crystalline solids at room temperature.
- Ionic compounds have high melting and boiling points.

Covalent Bonding

Atoms combine to form covalent molecules by sharing electrons. This process is called covalent bonding. The resulting molecule is more stable than the individual atoms.

Types of Covalent Bonds:

- Single Bond:** Formed by the sharing of one pair of electrons between two atoms.
- Double Bond:** Formed by the sharing of two pairs of electrons between two atoms.
- Triple Bond:** Formed by the sharing of three pairs of electrons between two atoms.

Properties of Covalent Bonds:

- Covalent bonds are strong and directional.
- Covalent molecules are gases, liquids, or solids at room temperature.
- Covalent molecules have lower melting and boiling points than ionic compounds.

A diagram showing two atoms, one with a positive charge (+) and one with a negative charge (-), interacting. The positive atom has a single electron in its outer shell, and the negative atom has a single electron in its outer shell. They are shown moving towards each other, with dashed lines indicating the path of the electrons. The final state shows the positive atom with two electrons in its outer shell and the negative atom with two electrons in its outer shell, forming a stable ionic bond.

GUIDED READING ACTIVITY

Ionic Bonds Guided Reading And Study Answers

Michael J. Padilla



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