

IONIC COVALENT BONDING WEBQUEST

IONIC BONDING

WEBSITE 1: <http://visionlearning.com/en/library/Chemistry/1/Chemical-Bonding/55>

1. Approximately how many elements are represented on the periodic table?
118
2. What accounts for the fact that there are far more substances than are listed on the periodic table?
there are more substances because elements can combine and react to form new substances
3. "Formed when two or more atoms chemically bond together, the resulting compound is unique both chemically and physically from its parent atoms."
4. What compound forms during the reaction between elemental sodium and elemental chlorine?
Name: sodium chloride
Formula: NaCl
5. List four facts related to G.N. Lewis:
 - atoms with fewer than eight electrons bond together
 - he is an american chemist
6. Explain the behavior of electrons in ionic bonding.
electrons are transferred from one atom to another
7. In the process of either gaining or losing negatively charged electrons, the reacting atoms form ions.
8. In the reaction between sodium and chlorine, which atom loses an electron?
Na loses e- Which atom gains an electron? Cl gains e-
9. After transferring the electron, which ion is negatively charged? The anion
Why is the ion negatively charged? it gains an electron Write the symbol for the negative ion: Anion(-)
10. After transferring the electron, which ion is positively charged? the cation ion
Why is the ion positively charged? it loses electrons
Write the symbol for the positive ion: C(+)
11. How does the sodium atom contrast to the sodium ion? (charge, size)
The sodium loses the electrons it becomes smaller in size and its charge become positive.

Ionic Covalent Bonding Webquest Answers Key Page

Kok Koon Leong



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Ionic and Covalent Bonding James M. Postma, Julian Roberts, J. Leland Hollenberg, 2000-01-15 *The Covalent Bond* Henry Sinclair Pickering, 1977 **Chemical Bonding** M.S. Sethi & M. Satake, 2010 Contents Chemical Bonding I Basic Concepts Chemical Bonding II Additional Aspects Intermolecular Force and Crystal Structures **Chemical Bonding** Kok Koon Leong, 2021-05-03 This book introduces the principles behind chemical bonding to teenagers between the ages of fifteen to seventeen Topics covered include ionic bonding covalent bonding and metallic bonding *Teaching Chemical Bonding* Margaret Irene Lindsay, 1995 This document presents an instructional strategy for teaching chemical bonding using parables and music Games student interactions and worksheets are included in the lesson plans Topics include metallic bonding covalent bonding including molecular and network structure and ionic bonding JRH

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