

# IONIC BOND

## Definition & Key Concepts



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### Definition

An ionic bond is a chemical bond formed due to the transfer of electrons from one atom to another, resulting in oppositely charged ions that attract each other.

♦ Example: NaCl (Sodium Chloride) - Table Salt

Ionic bonding concepts evolved from **Faraday's electrolysis experiments**, later refined by **Lewis and Kossel**, who proposed atoms transfer electrons to achieve stability.



### HOW DOES IT FORM?

- 1 Electron Transfer:** One atom donates an electron, another gains it.
- 2 Ion Formation:** The donor becomes a positively charged cation (e.g.,  $\text{Na}^+$ ), and the receiver becomes a negatively charged anion (e.g.,  $\text{Cl}^-$ ).
- 3 Electrostatic Attraction:** Oppositely charged ions attract, forming a strong bond.

### CHARACTERISTICS OF IONIC BOND

- **Electron Transfer:** Complete transfer of electrons.
- **Electrostatic Attraction:** Strong attraction between oppositely charged ions.
- **Ion Formation:** Forms cations (+) and anions (-).
- **Crystal Lattice:** Arranges in a rigid lattice structure.
- **High Melting & Boiling Points:** Due to strong ionic bonds.
- **Conductivity:** Conducts electricity in molten or dissolved state, not in solid.



### COMMON EXAMPLES OF IONIC COMPOUNDS

- 📖 Sodium Chloride ( $\text{NaCl}$ ) - Table Salt
- 🦷 Calcium Fluoride ( $\text{CaF}_2$ ) - Found in Toothpaste
- ⚡ Potassium Bromide ( $\text{KBr}$ ) - Used in Photography
- 💊 Magnesium Oxide ( $\text{MgO}$ ) - Found in Medicines

# Ionic Bond Review And Reinforce

**Rosina Ehmann**



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