

Unit 5 - Ionic Bonding Study Guide

Name: KEY Period: Date:

Review Ionic Bonds

- What two types of elements will transfer electrons to form an ionic bond?
Metal & nonmetal
- What is an ionic bond?
An electrostatic force that holds together a cation and anion
- How can you describe the electrical charge of an ionic compound?
Neutral
- Metals will lose all of their valence electrons to form a positively charged cation. The oxidation number will be equal to the number of valence electrons.
- Nonmetals will gain valence electrons to form a negatively charged anion. The oxidation number will be equal to 8 – the # of valence electrons.
- Which elements are least likely to undergo bonding? *Noble Gases*
- Write the oxidation number for the following elements.

a. Cesium	<u>+1</u>	e. Magnesium	<u>+2</u>
b. Potassium	<u>+1</u>	f. Barium	<u>+2</u>
c. Aluminum	<u>+3</u>	g. Nitrogen	<u>-3</u>
d. Iron (III)	<u>+3</u>	h. Iodine	<u>-1</u>
		i. Oxygen	<u>-2</u>
		j. Selenium	<u>-2</u>
- Which alkali metal has the highest ionization energy? *Hydrogen*
- Which alkaline earth metal has the lowest ionization energy? *Radium*
- What properties characterize ionic compounds?
 - High melting point & boiling point*
 - Solid at room temperature*
 - Brittle*
 - Ionic solids can't conduct electricity*
 - Molten ionic compounds can conduct electricity*

Transferring of Electrons

11. Please draw the transfer of electrons for the following:

a. NaBr



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M Woodhall



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