energy level

# CHAPTER 7 STUDY GUIDE

# Ionic Compounds and Metals

#### Section 7.1 Ion Formation

chemical bond

In your textbook, read about chemical bonds and formation of ions.

electrons

Use each of the terms below just once to complete the passage.

becoming anions.

nucieus	ociet	pseudo-noble gas formations	valuence
The force	that holds two atoms together	is called a(n) (1)	
Such an attac	hment may form by the attract	ion of the positively charged	
(2)	of one atom for the negatively charged		
(3)	of another atom, or by the attraction of charged atoms, which are called		
	. The attractions may also involve		
(5)	electrons, whic	th are the electrons in the outermost	
(6)	. The (7)	are a family of elements	that have very
little tendenc	y to react. Most of these eleme	nts have a set of eight outermost electrons,	which is called a
stable (8)	. The relatively stable electron structures developed by loss of		
electrons in c	ertain elements of groups 3, 4,	13, and 14 are called (9)	
For each sta	tement below, write true or fa	alse.	
	10. A positively charged ion is called an anion.		
	<ol> <li>Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.</li> </ol>		
	<ol><li>Elements tend to react so that they acquire the electron structure of a halogen.</li></ol>		
	<ol> <li>A sodium atom tends to lose one electron when it reacts.</li> </ol>		
	14. The electron structure of a zinc ion (Zn <sup>2+</sup> ) is an example of a pseudo-noble gas formation.		
	15. A Cl <sup>-</sup> ion is an examp	de of a cation.	
	16. The ending -ide is used to designate an anion.		
	17 Nonmostale form a sta-	ble outer electron configuration by losing a	destrone and

noble gases

BOOK

# **Ionic Compounds And Metals Study Guide**

**M** Lipman

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