

Unit #1

Chapter #13: Physical Properties of Solutions

- What are the Types of Solutions?

- What are the Parts of a Solution?

Solution - homogeneous mixture



- 2) The rubbing alcohol sold in drug stores often is composed of 70% isopropyl alcohol and 30% water. In this solution

- ☐ A) water is the solvent.
- ☒ B) isopropyl alcohol is the solvent.
- ☐ C) both water and isopropyl alcohol are solvents.
- ☐ D) neither water nor isopropyl alcohol is a solvent.



- Unsaturated, Saturated, Supersaturated?

Unsaturated → less than dissolving capacity

Saturated → reaching max dissolving amount

Supersaturated → More is dissolved than allowed (high temp cooled)

- 11) An unsaturated solution is one that _____

- ☐ A) has a concentration lower than the solubility
- ☐ B) has no double bonds
- ☐ C) contains the maximum concentration of solute possible, and is in equilibrium with undissolved solute
- ☒ D) contains no solute
- ☐ E) contains more dissolved solute than the solubility allows

saturated →
supersaturated →

- What are the Solution Units?

$$\text{Molarity (M)} = \frac{\text{moles of solute}}{\text{Volume of solution (L)}}$$

$$\text{Molality (m)} = \frac{\text{moles of solute}}{\text{mass of solute (kg)}}$$

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