CHAPTER 9 Chemical Reactions

Section 1 Reactions and Equations

pages 282-288

Practice Problems

pages 284-287

Write skeleton equations for the following word equations.

 Hydrogen and bromine gases react to yield hydrogen bromide.

 $hydrogen(g) + bromine(g) \rightarrow hydrogen bromide(g)$

 $H_2(g) + Br_2(g) \rightarrow HBr(g)$

When carbon monoxide and oxygen react, carbon dioxide forms.

carbon monoxide(g) + $oxygen(g) \rightarrow carbon dioxide(g)$

 $CO(g) + O_2(g) \rightarrow CO_2(g)$

 Challenge Write the word equation and the skeleton equation for the following reaction: when heated, solid potassium chlorate yields solid potassium chloride and oxygen gas.

Potassium chlorate(s) → potassium chloride(s) + oxygen(g)

 $KClO_3(s) \rightarrow KCl(s) + O_2(g)$

Write chemical equations for each of the following reactions.

 In water, iron(III) chloride reacts with sodium hydroxide, producing solid iron(III) hydroxide and sodium chloride.

 $FeCl_2(aq) + 3NaOH(aq) \rightarrow Fe(OH)_3(s) + 3NaCl(aq)$

Liquid carbon disulfide reacts with oxygen gas, producing carbon dioxide gas and sulfur dioxide gas.

 $CS_2(l) + 3O_2(g) \rightarrow CO_2(g) + 2SO_2(g)$

 Challenge A piece of zinc metal is added to a solution of dihydrogen sulfate. This reaction produces a gas and a solution of zinc sulfate.

 $Zn(s) + H_2SO_{\theta}(aq) \rightarrow H_2(g) + ZnSO_{\theta}(aq)$

Section 1 Review

page 288

Explain why it is important that a chemical equation be balanced.

Because mass is neither created nor destroyed in chemical reactions, the numbers of atoms of all elements must be equal on both sides of the reaction arrow.

 List three types of physical evidence that indicate a chemical reaction has occurred.

Answers may include release or absorption of energy, change in color, change in odor, formation of a gas, or formation of a solid.

Compare and contrast a skeleton equation and a chemical equation.

The skeleton equation includes the formulas of reactants and products. The chemical equation gives the relative amounts of reactants and products.

 Explain why it is important to reduce coefficients in a balanced equation to the lowest possible wholenumber ratio.

Coefficients in the lowest ratio most clearly indicate the relative amounts of substances in a reaction.

 Analyze When balancing a chemical equation, can you adjust the subscript in a formula? Explain.

No. Doing so changes the identity of the substance.

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Technical Committee on Hazardous
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NFPA 491M National Fire Protection Association. Technical Committee on Hazardous Chemical Reactions, American Manual of Hazardous Chemical Reactions National Fire Protection Association. National Standards Institute, 1991 Technical Committee on Hazardous Chemical Reactions, American National Standards Institute, 1991 Hazardous Chemical Reactions. A Compilation of Chemical Reactions Reported to be Potentially Hazardous National Fire Protection Association. Committee on Chemicals and Explosives. Sectional Committee on Hazardous Chemical Manual of Hazardous Chemical Reactions ,1971 Reactions, 1964 **NFPA** National Fire Protection Association, 1975 Manual of Hazardous Chemical Reactions National Fire Protection Association, Committee on Chemicals and Manual of Hazardous Chemical Reactions National Fire Protection Association, 1964 Manual of Explosives, 1975 Hazardous Chemical Reactions Robert W. van Dolah, National Fire Protection Association (United States), 1975 Manual of Hazardous Chemical Reactions National fire protection association. Committee on chemicals and explosives. Sectional committee on hazardous chemical reactions, 1975 Manual of Hazardous Chemical Reactions 1975 National Fire Protection Association, Sectional Committee on Hazardous Chemical Reactions, 1975 **Manual of Hazardous Chemical Reactions** National Fire Protection Association. Committee on Chemicals and Explosives, 1968 Manual of Hazardous Chemical Reactions 1966 National Fire Protection Association. 1966 Manual of Hazardous Chemical Reactions, 1968 **Manual of Hazardous Chemical Reactions**, 1968 Manual of Hazardous Chemical Reactions, 1968 ,1968 Manual of Hazardous Chemical Reactions National Fire Protection National Fire Protection Association (U.S.), 1968 Manual of Hazardous Chemical Reactions, 1986 National Fire Protection Association, 1986 Association. NTPA., 1966 Manual of Hazardous Chemical Reactions, 1975 National Fire Protection Association, 1975 Manual of Hazardous Chemical Reactions, 1971 National Fire Protection Association, 1971 Manual of Hazardous Chemical Reactions 1986 National Fire Prevention Association, Technical Committee on Hazardous Chemical Reactions, 1986

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