



EZGO Repair Manual

1970 to 1990

GX-440
&
GX-444

Gas Golf Carts

Introduction

This Operation and service manual has been prepared to serve two Purposes. First, it will acquaint the reader with the operation of the EZGO Product and assist him in performing minor service and re-pairs. Secondly, it will introduce to the mechanic the most efficient Service and specifications and procedures. It is our hope that these procedures together with the trouble-shooting Tips outlined herein will enable the customer to receive the best Possible service from our product.

Manual For Ez Go 1987 Model 830

Montgomery Ward



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web consider the thermochemical equation below

$$4\text{NO(g)} + 6\text{H}_2\text{O(l)} + 4\text{NH}_3\text{(g)} + 5\text{O}_2\text{(g)} \rightarrow 11\text{H}_2\text{O(l)} + 4\text{NO}_2\text{(g)}$$
 is this reaction endothermic or exothermic
 endothermic
 how much heat will be absorbed or released if 10 g of NO is reacted with 10 g of H_2O

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web heats of fusion and vaporization determines what determines how much energy is required to change a substance's phase
 what is the heat of fusion formula
 $q = n \cdot \Delta H_{\text{fus}}$
 what is the heat of vaporization formula
 $q = n \cdot \Delta H_{\text{vap}}$
 what is on the x axis of the heating curve
 time

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 3 carry out a thermochemical calculation on all these species at their optimized geometries
 4 sum up the enthalpies free energy of the products and the reagents keeping into account the

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web answer a 1 explanation the entire reaction was multiplied by 2 therefore ΔH must also be halved
 3 answer c explanation
 option c is incorrect because 566 kJ/mol is the ΔH per mole of O_2 and not for 2 moles even though options b and d are expressed in grams they are still correct because 56

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 what is the heat change when 472 g of carbon reacts with excess O_2 according to the following equation

$$\text{C(s)} + \text{O}_2\text{(g)} \rightarrow \text{CO}_2\text{(g)} \quad \Delta H = -393.5 \text{ kJ/mol}$$
 how much heat energy must be expelled by the body to rid the body of one pound of fat
 2s 3o 2 2so 3 h 7914 kJ is this endothermic or exothermic
 2

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 exothermic transfer of heat from the system to the surroundings
 $2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \rightarrow 2\text{H}_2\text{O(l)}$
 energy endothermic the transfer of heat from the surroundings to the system
 $2\text{HgO(s)} \rightarrow 2\text{Hg(l)} + \text{O}_2\text{(g)}$

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 enthalpy is an extensive property depends on amount of matter present this fact means that the magnitude of h is directly proportional to the amount of reactant consumed in the process

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web q 140 g 4 184 J 160°C g 0°C q 9372 J 16 J 35.5 g x 1 mole KCN 0.5453 moles 65.1 g 0.5453 mole moles 5
 when a 15 g sample of SrCl_2 dissolves in 80 g of water in a calorimeter the temperature drops from 42°C to 15°C
 calculate the ΔH for the process

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