

PHYSICAL KINETIC MOLECULAR THEORY

Learning Objectives

- 1) Identify theoretical differences in how gases participate in the kinetic theory of gases and how molecules in liquids and solids participate in the kinetic theory of gases and molecular theory.
- 2) Identify how temperature affects molecular motion.
- 3) Explain how pressure affects molecular motion and how it is related to molecular theory.

Module 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in gases, liquids, and solids.



Figure 1: Representation of particles in different phases of matter.

Understanding the differences in how particles behave in different phases is essential for understanding the kinetic theory of gases and molecular theory.

Key Concepts

1. **Gas:** Gas particles are widely spaced and move rapidly in all directions. They have the most freedom of movement and are not held together by strong forces.
2. **Liquid:** Liquid particles are in contact but can move past each other. They have less freedom of movement than gas particles but more than solid particles.
3. **Solid:** Solid particles are in contact and have a fixed position. They have the least freedom of movement and are held together by strong forces.

Table 1: Comparison of the Kinetic Theory of Gases.

	1. GAS	2. LIQUID	3. SOLID
Particle arrangement:	Particles are widely spaced and move rapidly in all directions. They have the most freedom of movement and are not held together by strong forces.	Particles are in contact but can move past each other. They have less freedom of movement than gas particles but more than solid particles.	Particles are in contact and have a fixed position. They have the least freedom of movement and are held together by strong forces.
Particle movement:	Particles move rapidly in all directions. They have the most freedom of movement and are not held together by strong forces.	Particles move past each other. They have less freedom of movement than gas particles but more than solid particles.	Particles vibrate in place. They have the least freedom of movement and are held together by strong forces.

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