Mole Calculations Review Worksheet - answers on next page.

Complete the problems on this worksheet to review for the Chapter 10 test. Be sure to use dimensional analysis, and label all numbers with units when solving the problems.

Calculate the molar mass of each compound.

a. LiOH

c. Mg(C₂H₂O₂)₃

b. barium bromide

d. Ca(NO_{d)}

- 2. How many molecules are in 45.0 grams CH₄?
- 3. How many moles are in 18.8 grams NaOH?
- 4. A salt shaker containing 9.58 x 10²³ formula units NaCl contains how many moles?
- 5. Which of the following has the greatest mass?

4.2 moles C

8.34 x 10³⁴ f.u. CaCO₃ or 12.6 g Al(NO₃)₃

Find the number of representative particles in each of the following:

a. 0.28 mol Ni

c. 1.32 mol F₂

b. 1.84 mol S

d. 3.15 mol K.PO.

- How many liters would 1.45 moles oxygen gas occupy at STP?
- How many grams would 15.8 L of ammonia occupy at STP?
- 9. How many total atoms are in 5.8 mol CaCl₂?
- 10. Calculate the percent composition of each element in the following compounds:

a. iron(III) oxide

c. methane

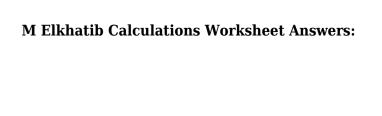
b. Ca₃(PO₂)₂

d. H₂SO₄

- 11. Calculate the empirical and molecular formulas of adrenaline. The compound has a molecular mass of 183 g/mol, and is 59.0% C, 7.1% H, 26.2% O, and 7.7% N by mass.
- 12. A compound is 72.2% Mg and 27.8% N by mass. Calculate the empirical formula.
- 13. Nicotine is 74.1% C, 8.6% H, and 17.3% N by mass. The molecular mass of nicotine is 162 g/mol. Calculate its empirical and molecular formulas.
- 14. How many total ions are in 2.95 grams MgSO₁?

M Elkhatib Calculations Worksheet Answers

Michael Seilmaier



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