

- 2 Some metal carbonates cannot be obtained in a pure state. For example magnesium carbonate exists in a 'basic' form, in which magnesium hydroxide is also present.

One possible chemical formula of basic magnesium carbonate is $\text{MgCO}_3 \cdot \text{Mg}(\text{OH})_2 \cdot 2\text{H}_2\text{O}$.

When basic magnesium carbonate is heated, if the possible formula were correct, it would decompose as shown below.



In this experiment, you will decompose basic magnesium carbonate by heating it, and you will use your results to determine whether this possible formula is correct.

FB 4 is basic magnesium carbonate.

(a) Method

Read through the method before starting any practical work and prepare a table for your results in the space below.

- Weigh a crucible with its lid and record the mass.
- Add 1.1-1.3g of **FB 4** to the crucible. Weigh the crucible and lid with **FB 4** and record the mass.
- Place the crucible on the pipe-clay triangle and remove the lid.
- Heat the crucible and contents **gently** for about one minute.
- Then heat the crucible and contents strongly for about four minutes.
- Replace the lid and allow the crucible to cool for at least five minutes.
- **While the crucible is cooling, you may wish to begin work on Question 3.**
- Re-weigh the crucible and contents with lid. Record the mass.
- Calculate, and record, the mass of **FB 4** used and the mass of residue obtained.

Isa Paper 2 Chemistry Metal Carbonates Ore

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Table of Contents Isa Paper 2 Chemistry Metal Carbonates Ore

1. Understanding the eBook Isa Paper 2 Chemistry Metal Carbonates Ore
 - The Rise of Digital Reading Isa Paper 2 Chemistry Metal Carbonates Ore
 - Advantages of eBooks Over Traditional Books
2. Identifying Isa Paper 2 Chemistry Metal Carbonates Ore
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Isa Paper 2 Chemistry Metal Carbonates Ore
 - User-Friendly Interface
4. Exploring eBook Recommendations from Isa Paper 2 Chemistry Metal Carbonates Ore
 - Personalized Recommendations
 - Isa Paper 2 Chemistry Metal Carbonates Ore User Reviews and Ratings
 - Isa Paper 2 Chemistry Metal Carbonates Ore and Bestseller Lists

5. Accessing Isa Paper 2 Chemistry Metal Carbonates Ore Free and Paid eBooks
 - Isa Paper 2 Chemistry Metal Carbonates Ore Public Domain eBooks
 - Isa Paper 2 Chemistry Metal Carbonates Ore eBook Subscription Services
 - Isa Paper 2 Chemistry Metal Carbonates Ore Budget-Friendly Options
6. Navigating Isa Paper 2 Chemistry Metal Carbonates Ore eBook Formats
 - ePub, PDF, MOBI, and More
 - Isa Paper 2 Chemistry Metal Carbonates Ore Compatibility with Devices
 - Isa Paper 2 Chemistry Metal Carbonates Ore Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Isa Paper 2 Chemistry Metal Carbonates Ore
 - Highlighting and Note-Taking Isa Paper 2 Chemistry Metal Carbonates Ore
 - Interactive Elements Isa Paper 2 Chemistry Metal Carbonates Ore
8. Staying Engaged with Isa Paper 2 Chemistry Metal Carbonates Ore
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Isa Paper 2 Chemistry Metal Carbonates Ore
9. Balancing eBooks and Physical Books Isa Paper 2 Chemistry Metal Carbonates Ore
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Isa Paper 2 Chemistry Metal Carbonates Ore
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Isa Paper 2 Chemistry Metal Carbonates Ore
 - Setting Reading Goals Isa Paper 2 Chemistry Metal Carbonates Ore
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Isa Paper 2 Chemistry Metal Carbonates Ore
 - Fact-Checking eBook Content of Isa Paper 2 Chemistry Metal Carbonates Ore
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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