

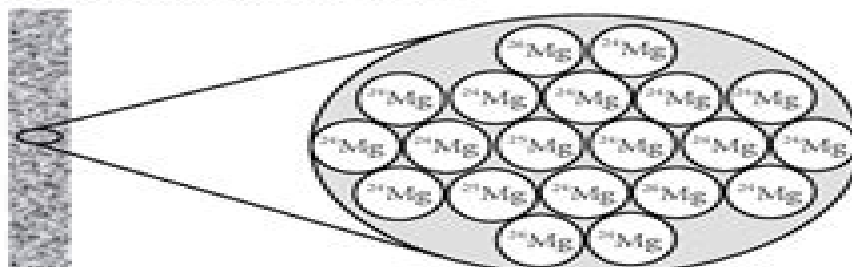
## Average Atomic Mass

How are the masses on the periodic table determined?

### Why?

Most elements have more than one naturally occurring isotope. As you learned previously, the atoms of those isotopes have the same atomic number (number of protons), making them belong to the same element, but they have different mass numbers (total number of protons and neutrons) giving them different atomic masses. So which mass is put on the periodic table for each element? Is it the most common isotope's mass? The heaviest mass? This activity will help answer that question.

### Model 1 – A Strip of Magnesium Metal



1. Write in the atomic number for each Mg atom in Model 1.

All are 12

2. What are the mass numbers of the naturally occurring isotopes of magnesium shown in Model 1?

24, 25, 26

3. Do all of the atoms of magnesium in Model 1 have the same atomic mass? Explain.

No there are 3 different mass numbers which mean there are 3 different masses

4. For the sample of 20 atoms of magnesium shown in Model 1, draw a table indicating the mass numbers of the three isotopes and the number of atoms of each isotope present.

Mass #	# of atoms
24	16
25	2
26	2

5. Which isotope of magnesium is the most common in Model 1?

Magnesium 24 is the most common

6. Based on Model 1 and the table you created in Question 4, for every 10 atoms of magnesium, approximately how many atoms of each isotope will be found?

Mg 24=8 Mg 25=1 Mg 26=1

# Isotopes And Average Atomic Mass Answer Key

**Daniel F McAuley**



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## **Table of Contents Isotopes And Average Atomic Mass Answer Key**

1. Understanding the eBook Isotopes And Average Atomic Mass Answer Key
  - The Rise of Digital Reading Isotopes And Average Atomic Mass Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Isotopes And Average Atomic Mass Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Isotopes And Average Atomic Mass Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Isotopes And Average Atomic Mass Answer Key
  - Personalized Recommendations
  - Isotopes And Average Atomic Mass Answer Key User Reviews and Ratings
  - Isotopes And Average Atomic Mass Answer Key and Bestseller Lists
5. Accessing Isotopes And Average Atomic Mass Answer Key Free and Paid eBooks
  - Isotopes And Average Atomic Mass Answer Key Public Domain eBooks
  - Isotopes And Average Atomic Mass Answer Key eBook Subscription Services
  - Isotopes And Average Atomic Mass Answer Key Budget-Friendly Options

6. Navigating Isotopes And Average Atomic Mass Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Isotopes And Average Atomic Mass Answer Key Compatibility with Devices
  - Isotopes And Average Atomic Mass Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Isotopes And Average Atomic Mass Answer Key
  - Highlighting and Note-Taking Isotopes And Average Atomic Mass Answer Key
  - Interactive Elements Isotopes And Average Atomic Mass Answer Key
8. Staying Engaged with Isotopes And Average Atomic Mass Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Isotopes And Average Atomic Mass Answer Key
9. Balancing eBooks and Physical Books Isotopes And Average Atomic Mass Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Isotopes And Average Atomic Mass Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Isotopes And Average Atomic Mass Answer Key
  - Setting Reading Goals Isotopes And Average Atomic Mass Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Isotopes And Average Atomic Mass Answer Key
  - Fact-Checking eBook Content of Isotopes And Average Atomic Mass Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

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