

Experiment 5

Determination of the Percentage of Ligands in $K_3[Fe(C_2O_4)_3] \cdot 3H_2O$

OBJECTIVES

1. To standardise the potassium permanganate solution.
2. To determine the percentage of ligands in $K_3[Fe(C_2O_4)_3] \cdot 3H_2O$.

INTRODUCTION

Standardisation of potassium permanganate ($KMnO_4$) solution is carried out by titration against a standard solution of oxalic acid. Oxalic acid is a good primary standard because the compound is available in solid form, as $H_2C_2O_4 \cdot 2H_2O$, which can be prepared to a very high degree of purity and is not hygroscopic nor efflorescent.

The reaction is as follows:



OVERALL REACTION



The following points about this reaction should be noted:

1. The reaction occurs only in acidic solutions. Acidification is achieved by the addition of sulfuric acid.
2. The reaction rate is normally slow. However, the reaction rate in this experiment is increased by (i) warming the contents of the conical flask or (ii) catalysed by the presence of manganese(II) ions, which is a product of the reaction.

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Marco Cascella



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Table of Contents Lab Manual Experiment 8 Wiley

1. Understanding the eBook Lab Manual Experiment 8 Wiley
 - The Rise of Digital Reading Lab Manual Experiment 8 Wiley
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Manual Experiment 8 Wiley
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Manual Experiment 8 Wiley
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Manual Experiment 8 Wiley
 - Personalized Recommendations
 - Lab Manual Experiment 8 Wiley User Reviews and Ratings
 - Lab Manual Experiment 8 Wiley and Bestseller Lists

5. Accessing Lab Manual Experiment 8 Wiley Free and Paid eBooks
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 - Lab Manual Experiment 8 Wiley Budget-Friendly Options
6. Navigating Lab Manual Experiment 8 Wiley eBook Formats
 - ePub, PDF, MOBI, and More
 - Lab Manual Experiment 8 Wiley Compatibility with Devices
 - Lab Manual Experiment 8 Wiley Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Manual Experiment 8 Wiley
 - Highlighting and Note-Taking Lab Manual Experiment 8 Wiley
 - Interactive Elements Lab Manual Experiment 8 Wiley
8. Staying Engaged with Lab Manual Experiment 8 Wiley
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Manual Experiment 8 Wiley
9. Balancing eBooks and Physical Books Lab Manual Experiment 8 Wiley
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Manual Experiment 8 Wiley
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lab Manual Experiment 8 Wiley
 - Setting Reading Goals Lab Manual Experiment 8 Wiley
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lab Manual Experiment 8 Wiley
 - Fact-Checking eBook Content of Lab Manual Experiment 8 Wiley
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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